

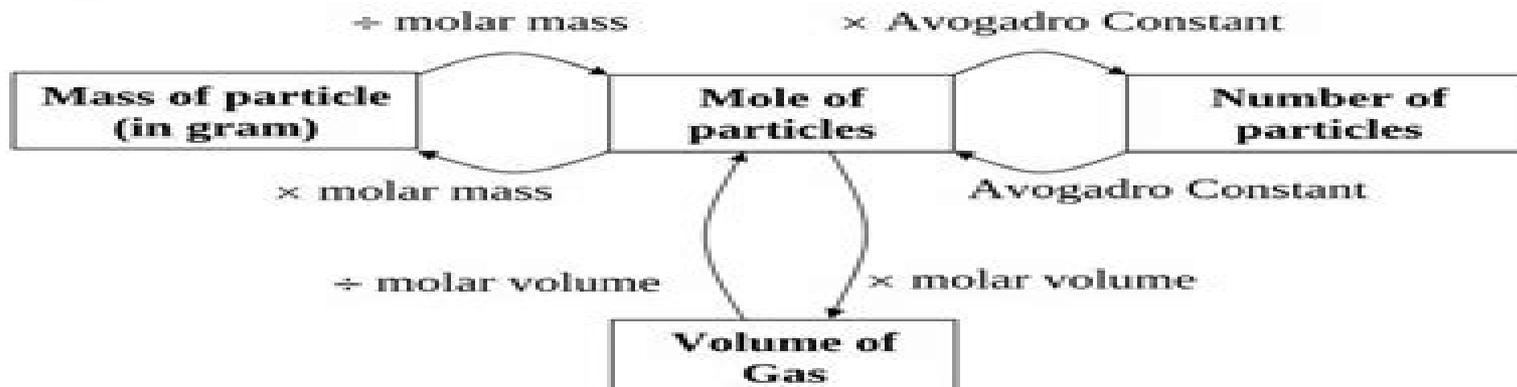
Short Notes: Form 4 Chemistry

Chemical Formulae and Equation

Calculation

<p>For Solid, liquid or gas</p> $\text{number of mole} = \frac{\text{mass of substance}}{\text{molar mass}}$ <p>Molar mass = RAM/RMM/RFM in gram</p>	<p>For gas (only)</p> $\text{number of mole} = \frac{\text{volume of gas}}{\text{molar volume}}$ <p>Molar volume = 24dm³ at room temperature Molar volume = 22.4dm³ at s.t.p.</p>
<p>For Solution</p> $\text{number of mole} = \frac{MV}{1000}$ <p>M = molarity V = Volume of solution in cm³</p>	<p>For quantity of particle(atom,molecule,ion)</p> $\text{number of mole} = \frac{\text{quantity of particle}}{6.02 \times 10^{23}}$

Summary



One School Short Notes Form 4 Chemistry

D Siedentop



One School Short Notes Form 4 Chemistry:

One School Short Notes Form 4 Chemistry Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**One School Short Notes Form 4 Chemistry**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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