

$$K_c = \frac{0.6 + x}{(0.2 - x)(0.3 - x)} = 30$$

$$0.6 + x = 30(0.06 - 0.2x - 0.3x + x^2)$$

$$0.6 + x = 1.8 - 15x + 30x^2$$

$$30x^2 - 16x + 1.2 = 0$$

$$x_{1,2} = \frac{16 \pm \sqrt{256 - 144}}{60} = \frac{16 \pm 10.6}{60}$$

$$x_1 = 0.44 \quad x_2 = 0.09$$

$$[SO_2] = 0.2 - 0.09 = 0.011 \text{ mol/l}$$

$$[Cl_2] = 0.3 - 0.09 = 0.021 \text{ mol/l}$$

$$[SO_2 Cl_2] = 0.6 - 0.09 = 0.69 \text{ mol/l}$$

3) Per il seguente equilibrio in fase omogenea la  $K_c$  vale 5.06:



Sapendo che ad un certo stadio del processo le concentrazioni istantanee di CO, H<sub>2</sub>O, CO<sub>2</sub>, e H<sub>2</sub> sono rispettivamente 0.1 M, 0.2 M, 1 M e 0.5M, stabilire se la reazione è all'equilibrio e, in caso contrario, dire in che senso avverrà e calcolare la composizione all'equilibrio (si consideri il volume di 1 litro).

	CO	+	H <sub>2</sub> O	⇌	CO <sub>2</sub>	+	H <sub>2</sub>
inizio	0.1		0.2		1		0.5

$$Q = \frac{[CO_2][H_2]}{[CO][H_2O]} = \frac{1 \cdot 0.5}{0.1 \cdot 0.2} = 25 \quad Q > K_c \quad \text{quindi la reazione è spostata verso sinistra}$$



	CO	+	H <sub>2</sub> O	⇌	CO <sub>2</sub>	+	H <sub>2</sub>
inizio	0.1		0.2		1		0.5
equilibrio	(0.1+x)		(0.2+x)		(1-x)		(0.5-x)

$$K_c = \frac{[CO_2][H_2]}{[CO][H_2O]}$$

$$5.06 = \frac{(1-x)(0.5-x)}{(0.1+x)(0.2+x)} \quad 5.06 = \frac{0.5-x-0.5x+x^2}{0.02+0.1x+0.2x+x^2} \quad 5.06 = \frac{0.5-1.5x+x^2}{0.02+0.3x+x^2}$$

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**DP Hallahan**



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