

# Stoichiometry & the mole notes

## the mole concept

the mole (mol) is the SI unit for measuring the amount of a substance.

avogadro's number:

$$1 \text{ mol} = 6.022 \times 10^{23} \text{ particles}$$

(particles can be atoms, molecules, or formula units)

molar mass:

the mass of one mole of a substance in grams.

found by adding the atomic masses from the periodic table

example

$$\text{H}_2\text{O} (2 \cdot 1.008 \text{ g/mol}) + (1 \cdot 16.00 \text{ g/mol}) = 18.016 \text{ g/mol}$$

## Example Problems

1. how many atoms are in 3.5 moles of Fe?

$$3.5 \text{ mol Fe} \cdot \frac{6.022 \cdot 10^{23} \text{ atoms}}{1 \text{ mol}} = 2.11 \cdot 10^{24} \text{ atoms}$$

2. what is the mass of 2.00 moles of NaCl?

find molar mass:

$$\text{Na} : 22.99 \text{ g/mol}$$

$$\text{Cl} : 35.45 \text{ g/mol}$$

$$\text{Total} : 58.44 \text{ g/mol}$$

convert

$$2.00 \text{ mol} \cdot \frac{58.44 \text{ g}}{1 \text{ mol}} = 116.88 \text{ g}$$



## Conversions Involving the mole

conversion

particles  $\leftrightarrow$  moles

moles  $\leftrightarrow$  mass

moles  $\leftrightarrow$  volume (L) (at STP)

factor to use

$$1 \text{ mol} = 6.022 \cdot 10^{23} \text{ particles}$$

$$1 \text{ mol} = \text{molar mass (g)}$$

$$1 \text{ mol} = 22.4 \text{ L of gas}$$

## Stoichiometry: the study of quantitative relationships in chemical reactions

Balanced chemical equations show the mole ratio of reactants & products

mole ratios come from coefficients in a balanced equation.

Stoichiometry steps:

1. Balance the equation

2. convert given quantity into moles (if necessary)

3. Use mole ratio from balanced equation.

4. Convert moles back to desired unit (grams, liters, etc.)

Example:



question: how many grams of  $\text{H}_2\text{O}$  are produced from 5.00 moles of

$\text{O}_2$ ? Step 1: Use mole ratio

$$5.00 \text{ mol O}_2 \cdot \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol O}_2} = 10.0 \text{ mol H}_2\text{O}$$

Step 2: Convert to grams

$$10.0 \text{ mol H}_2\text{O} \cdot \frac{18.016 \text{ g}}{1 \text{ mol}} = 180.16 \text{ g H}_2\text{O}$$

## Limiting Reactant & Percent Yield

limiting reactant (LR): the reactant that runs out first, limiting product formation

excess reactant: the reactant that remains after the reaction is complete

Steps to identify LR

1. Convert reactants to moles

2. Use mole ratios to do

# Chemistry Notes Chapter 9 Stoichiometry

**Robert K. Wismer**



## Chemistry Notes Chapter 9 Stoichiometry:

*Chemistry of Pyrotechnics* Chris Mocella, John A. Conkling, 2019-01-15 This book provides chemists with technical insight on pyrotechnics and explosives It emphasizes basic chemical principles and practical hands on knowledge in the preparation of energetic materials It examines the interactions between and adaptations of pyrotechnics to changing technology in areas such as obscuration science and low signature flame emission The updated third edition discusses chemical and pyrotechnic principles components of high energy materials elements of ignition propagation and sensitivity It offers heat compositions including ignition mixes delays thermites and propellants and investigates the production of smoke and sound as well as light and color *Lecture-notes on Theoretical Chemistry* Ferdinand Gerhard Wiechmann, 1893 [High School Chemistry](#)

Michael A. Buben, 2018-05-18 The Teacher s Edition is designed to add direction to High School Chemistry which is an outline of notes The books go hand in hand Modified Second Edition as a student s workbook and guidance from the Teacher s Edition **Foundations of College Chemistry** Morris Hein, Susan Arena, 2013-01-01 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years Foundations of College Chemistry Alternate 14th Edition has helped readers master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis **Addison-Wesley**

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*Modern Chemistry* Nicholas D. Tzimopoulos, 1993 *Chemistry, Study Guide* Bernice G. Segal, Peter S.

Shenkin, 1989-02-14 This Second Edition of the first year chemistry text known for its clarity of exposition and its large number of illustrative worked problems contains a more rigorous treatment of electrochemistry chemical equilibrium and thermochemistry Worked examples now number over 300 and exercises over 1460 *Fundamentals of Ionic Liquids*

Douglas R. MacFarlane, Mega Kar, Jennifer M. Pringle, 2017-08-02 Written by experts who have been part of this field since its

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**Chemical Engineering Catalog** ,1922      **Prentice Hall Chemistry** ,2000      **High School Chemistry Teachers Magazine** ,1967

## **Chemistry Notes Chapter 9 Stoichiometry** Book Review: Unveiling the Magic of Language

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