

Empirical and Molecular Formulas Worksheet

Key

Objectives:

- be able to calculate empirical and molecular formulas

Empirical Formula

1) What is the empirical formula of a compound that contains 0.783g of Carbon, 0.196g of Hydrogen and 0.521g of Oxygen?

$$\begin{array}{l}
 \text{C} \quad \frac{0.783\text{g C}}{12.01\text{g}} = \frac{0.065}{0.033} \approx 1.96 \quad 2 \\
 \text{H} \quad \frac{0.196\text{g H}}{1.01\text{g}} = \frac{0.194}{0.033} \approx 5.88 \quad 6 \\
 \text{O} \quad \frac{0.521\text{g O}}{16.00\text{g}} = \frac{0.033}{0.033} = 1 \quad 1
 \end{array}$$

$\text{C}_2\text{H}_6\text{O}$

2) What is empirical formula of a compound which consists of 89.14% Au and 10.80% of O?

$$\begin{array}{l}
 \text{Au} \quad \frac{89.14\text{g}}{196.97} = \frac{0.45}{0.45} = 1 \quad \text{AuO}_{1.5} \\
 \text{O} \quad \frac{10.80\text{g}}{16\text{g}} = \frac{0.675}{0.45} = 1.5
 \end{array}$$

Au_2O_3

3) What is empirical formula if compound consists of 21.2%N, 6.1%H, 24.2%S and 48.5%O?

$$\begin{array}{l}
 \text{N} \quad \frac{21.2\text{g}}{14.01} = 1.5 \quad \cdot \quad 2 \\
 \text{H} \quad \frac{6.1\text{g}}{1.01\text{g}} = 6.03 \quad \cdot \quad 8 \\
 \text{S} \quad \frac{24.2\text{g}}{32} = 0.75 \quad \cdot \quad 1 \\
 \text{O} \quad \frac{48.5\text{g}}{16.00\text{g}} = 3.03 \quad \cdot \quad 4
 \end{array}$$

$\text{N}_2\text{H}_8\text{S}_4\text{O}_4$

Molecular Formula

4) Empirical formula of a substance is CH_2O . Molar mass is 180. What is the molecular formula?

$$\text{EFM} = (12 + 2 \times 0 + 16) = 30$$

$$\frac{180}{30} = 6 \quad \text{C}_6\text{H}_{12}\text{O}_6$$

5) Sample (3.585g) contains 1.388g of C, 0.345g of H, 1.850g O and its molar mass is 62g. What is molecular formula of this substance?

$$\begin{array}{l}
 \text{C} \quad \frac{1.388}{3.585} \rightarrow 38.7\% \rightarrow \frac{3.87\text{g}}{12.01\text{g}} = \frac{3.22}{3.22} = 1 \\
 \text{H} \quad \frac{0.345}{3.585} \rightarrow \frac{9.62\text{g}}{1.01\text{g}} = \frac{9.52}{3.22} = 3 \\
 \text{O} \quad \frac{1.850}{3.585} \rightarrow \frac{51.6\text{g}}{16\text{g}} = \frac{3.22}{3.22} = 1
 \end{array}$$

$\text{C}_2\text{H}_6\text{O}_2$

$\text{CH}_3\text{O} = \text{EFM} (12 + 3 \times 0 + 16)$

$$\frac{62}{31.03} = 2$$

1	$\text{C}_2\text{H}_6\text{O}$
2	Au_2O_3
3	$\text{N}_2\text{H}_8\text{S}_4\text{O}_4$
4	$\text{C}_2\text{H}_6\text{O}_2$
5	$\text{C}_2\text{H}_6\text{O}$

Chemistry Empirical Formula Worksheet Answers

Crispin Collins



Chemistry Empirical Formula Worksheet Answers:

Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998

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Mark Jackson,2006-02-22 This 6 page study guide contains basic chemistry analysis and concepts designed specifically to aid science students Chemical Calculations James Scott Long,Harold Victor Anderson,1928 Percent Composition and Empirical Formulas ,2016 In this lecture you encounter two classic types of chemistry problems and learn the basic characteristics of each The lecture concludes with several practice problems to help you master the skill of solving percent composition problems

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