

## Dimensional Analysis Practice Worksheet Solutions

Complete this and place it in your portfolio after section 2-2. I will email the solutions later so that you can check your work (make sure you are showing the work the same way you see it in the solutions, using dimensional analysis). It is important that you practice and learn to use dimensional analysis (multiplying by unit fractions) appropriately. Make sure you get help if you are still having difficulties with this topic. Use my office hours or the math tutors in the Student Learning Center in the library. You do not have to write sentences on this practice worksheet. You do need to show all of the units and use  $\approx$  for approximations.

Do the following conversions using dimensional analysis using the conversion factors on page 74 of your text. You should be doing these all in one computation (even if you have to use more than one unit fraction), not several computations. I have done #1 as example.

1. 5400 inches to miles

$$5400 \cancel{\text{in.}} \times \frac{1 \cancel{\text{ft}}}{12 \cancel{\text{in.}}} \times \frac{1 \text{ mi}}{5280 \cancel{\text{ft}}} = \frac{5400}{63360} \text{ mi} \\ \approx 0.085 \text{ mi}$$

2. 13 weeks to seconds

$$13 \cancel{\text{ weeks}} \times \frac{7 \cancel{\text{ days}}}{1 \cancel{\text{ week}}} \times \frac{24 \cancel{\text{ hrs}}}{1 \cancel{\text{ day}}} \times \frac{60 \cancel{\text{ min}}}{1 \cancel{\text{ hr}}} \times \frac{60 \text{ sec}}{1 \cancel{\text{ min}}} = 7,862,400 \text{ sec}$$

3. 54 yards to mm

$$54 \cancel{\text{ yards}} \times \frac{3 \cancel{\text{ feet}}}{1 \cancel{\text{ yd}}} \times \frac{1 \cancel{\text{ m}}}{3.28 \cancel{\text{ ft}}} \times \frac{1000 \text{ mm}}{1 \cancel{\text{ m}}} = \frac{162000}{3.28} \text{ mm} \\ \approx 49,390.24 \text{ mm}$$

4.  $36 \frac{\text{cm}}{\text{sec}}$  to mph

$$\frac{36 \cancel{\text{ cm}}}{1 \cancel{\text{ sec}}} \times \frac{60 \cancel{\text{ sec}}}{1 \text{ hr}} \times \frac{1 \cancel{\text{ m}}}{1000 \cancel{\text{ cm}}} \times \frac{3.28 \cancel{\text{ ft}}}{1 \cancel{\text{ m}}} \times \frac{1 \text{ mi}}{5280 \cancel{\text{ ft}}} = \frac{7084.8 \text{ mi}}{5280000 \text{ hr}} \\ \approx 0.00134 \frac{\text{mi}}{\text{hr}}$$

# Chemistry Dimensional Analysis Practice Problems Answers

**Yi-Tong Ma**



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topic to help calibrate your study plan

**Chemistry: Matter & Change, Solving Problems - A Chemistry Handbook** McGraw Hill, 2001-08  
**Glencoe Chemistry Solving Problems A Chemistry Handbook Matter and Change** Prentice Hall  
**Chemistry**, 2000 *Foundations of College Chemistry* Morris Hein, 1991 Features strong problem solving and carefully constructed problem sets

**Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl** Paul B. Kelter, 1993

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**General Chemistry** James P. Birk, Donald M. Baird, 1994-03 Preparatory Chemistry Howard Stephen Stoker, 1990  
Dosage and Solutions Calculations Alice J. Smith, 1989 Chemistry George Tyler Miller, 1987

**General Chemistry, Reactions First** Kevin Revell, 2024-12-04 Revell's General Chemistry empowers students to grasp essential topics and concepts with more ease Using a friendly approach the text uses metaphors and relatable examples to demystify even the most challenging subjects in general chemistry

Understanding Chemistry Fred M. Dewey, 1994 *Chemistry, Student Solutions Manual* James E. Brady, Joel W. Russell, John R. Holum, 2000-01-26

The images on the cover call attention to the relationship between macro observations and the intimate structure of chemical substances and the changes both chemical and physical that they undergo

**Fireworks** One of the ingredients is phosphorus a molecular form of which is believed to consist of linked tetrahedra of phosphorus atoms The chemical reaction of phosphorus with oxygen is partly responsible for the spectacular show of light

**Carbon** The element is found in several forms including the familiar diamond and another recently discovered sooty substance that consists of soccer ball shaped molecules often referred to as buckeyballs Diamond is not the most stable form of carbon and is created from other forms of carbon at high temperatures and pressures deep within the earth

**Acetylene torch** Cutting steel is possible because of the intense heat generated by the chemical reaction of acetylene with oxygen a reaction between molecules of  $C_2H_2$  and  $O_2$  to give  $CO_2$  and  $H_2O$

**Hot air balloon** The air that helps it rise is heated by the combustion of molecules of propane each composed of three carbon and eight hydrogen atoms

**Stormy weather** The evaporation of water serves to store energy provided by the sun Subsequent condensation of the water vapor releases this energy and is the basis of all the weather systems on our planet

Chemistry Theodore Lawrence Brown, 1988

**Chemistry** John W. Moore, Conrad L. Stanitski, Peter C. Jurs, 2002

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Gunter,2002-06 **Selected Titles in Chemistry** ,1972

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