

Chapter 1 - Introduction

Classifications of Matter – 1.3

- **Matter:** anything that occupies space and has mass
- **Chemistry:** the study of matter and its change
- **Atom:** smallest distinctive unit
- **Molecule:** two or more atoms together
- **Substance:** group of identical atoms or molecules
 - **Element:** atoms are identical
 - **Compound:** contains different atoms
- **Homogenous Mixture:** uniform, seems as one thing
- **Heterogeneous Mixture:** non-uniform, able to see the elements



Physical and Chemical Properties of Matter – 1.4

- **Physical Property:** measurable, describes system state
 - Melting point, solubility, density, color
 - "Air is a gas," "boiling point of water is 100 degrees," "Sulfur crystals are yellow"
- **Chemical Property:** chemical reaction, compound cannot be restored without a change
 - Burning, explosions, decomposition, rusting
 - i.e. "Burning Hydrogen in Oxygen to form water," "milk turning sour"



Measurement – 1.5

- **SI base units of Measurement:**
 - Length – meter (m)
 - Mass – kilogram (kg)
 - Time – second (s)
 - Temperature – Kelvin (K)
 - Amount of a substance – mole (mol)
- **Temperature Scales:**
 - Kelvin: absolute temperature scale, $273^{\circ}\text{C} = 1\text{K}$

Handling Number – 1.6

- **Accuracy:** Correctness – How close the results were to where you intended them to be
- **Precision:** Exactness – How close the results were in relation to other results / trials
- **Scientific Notation:** Expresses a number as a product of a number between 1 and 10 and the appropriate power of 10
 - All non-zeros are **significant** (1-9)
 - A zero between non-zeros is **significant** (209)
 - Final zeros to the right of the decimal are **significant** (10.0)
 - Initial zeros are **not significant** (0.47)
 - Final zeros in a number with no decimal **may or may not be significant** (9,500)
- **Adding and subtracting:** same number of digits past the decimal point as the number used with the least precision – the least amount of significant figures
 - If greater than 5, round up
- **Multiplication and Division:** least number of significant figures
- **Exact Numbers:** have no effect on significant figures – i.e. 60 seconds in a minute
- **Volume (V):** Length cubed, $\text{mL} = 1\text{cm}^3$

Chemistry 101 Study Guide

Patrick Vollmar



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