

Chapter 2 - Practice Test #1

7. An equilibrium system shifts left when the
- rate of the forward reaction is equal to the rate of the reverse reaction.
 - rate of the forward reaction is less than the rate of the reverse reaction.
 - rate of the forward reaction is greater than the rate of the reverse reaction.
 - rate of the forward reaction and the reverse reaction are constant.
8. A 1.00 L flask contains a gaseous equilibrium system. The addition of reactants to this flask results in a
- shift left and a decrease in the concentration of the products.
 - shift left and a increase in the concentration of the products.
 - shift right and a decrease in the concentration of the products.
 - shift right and a increase in the concentration of the products.
9. Consider the following equilibrium: $\text{CH}_4\text{(g)} + \text{H}_2\text{O(g)} + \text{heat} = \text{CO(g)} + 3\text{H}_2\text{(g)}$
In which of the following will both stresses shift the equilibrium to the right?
- a decrease in temperature and a decrease in volume
 - a increase in temperature and a decrease in volume
 - a decrease in temperature and a increase in volume
 - a increase in temperature and a increase in volume
10. Consider the following equilibrium: $2\text{SO}_2\text{(g)} + \text{O}_2\text{(g)} = 2\text{SO}_3\text{(g)}$ $\Delta H = -198 \text{ kJ}$
There will be no shift in this equilibrium when
- | | |
|--------------------------------|----------------------------------|
| A. more O_2 is added. | C. the volume is increased. |
| B. a catalyst is added. | D. the temperature is increased. |
11. Consider the following equilibrium: $2\text{Fe(s)} + 3\text{H}_2\text{O(g)} = \text{Fe}_2\text{O}_3\text{(s)} + 3\text{H}_2\text{(g)}$
The equilibrium expression is
- | | |
|---|--|
| A. $\text{Keq} = \frac{[\text{Fe}_2\text{O}_3][\text{H}_2]^3}{[\text{Fe}]^2[\text{H}_2\text{O}]^3}$ | B. $\text{Keq} = \frac{[\text{Fe}_2\text{O}_3][3\text{H}_2]}{[2\text{Fe}][3\text{H}_2\text{O}]}$ |
| C. $\text{Keq} = \frac{[\text{H}_2]^3}{[\text{H}_2\text{O}]^3}$ | D. $\text{Keq} = [\text{H}_2]^3$ |
12. Consider the following equilibrium: $\text{N}_2\text{O}_4\text{(g)} = 2\text{NO}_2\text{(g)}$ $\text{Keq} = 0.133$
At equilibrium, the $[\text{N}_2\text{O}_4]$ is equal to

Chemical Equilibrium Practice Test Answers

William Crowell Bray



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Table of Contents Chemical Equilibrium Practice Test Answers

1. Understanding the eBook Chemical Equilibrium Practice Test Answers
 - The Rise of Digital Reading Chemical Equilibrium Practice Test Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Equilibrium Practice Test Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemical Equilibrium Practice Test Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Equilibrium Practice Test Answers
 - Personalized Recommendations
 - Chemical Equilibrium Practice Test Answers User Reviews and Ratings
 - Chemical Equilibrium Practice Test Answers and Bestseller Lists
5. Accessing Chemical Equilibrium Practice Test Answers Free and Paid eBooks
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 - Chemical Equilibrium Practice Test Answers Enhanced eBook Features
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 - Highlighting and Note-Taking Chemical Equilibrium Practice Test Answers
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 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemical Equilibrium Practice Test Answers
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemical Equilibrium Practice Test Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Setting Reading Goals Chemical Equilibrium Practice Test Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemical Equilibrium Practice Test Answers
 - Fact-Checking eBook Content of Chemical Equilibrium Practice Test Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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