



COVALENT BONDING

Practice Problems

Do your best to solve each of the following problems.

SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following compounds as ionic or covalent.
 - NaCl
 - H_2O
 - HCl
 - KNO_3
 - H_2SO_4
 - CaCl_2
 - H_2O_2
 - H_2S
 - H_2SO_3
 - H_2SO_4
 - H_2SO_5
 - H_2SO_6
- Write the formula for the following compounds:
 - hydrogen peroxide
 - hydrogen sulfide
 - hydrogen cyanide
 - hydrogen selenide
 - hydrogen telluride
 - hydrogen selenate
 - hydrogen tellurate
 - hydrogen selenate(IV)
 - hydrogen tellurate(IV)
 - hydrogen selenate(VI)
 - hydrogen tellurate(VI)
- Write the name for each of the following compounds:
 - H_2O
 - H_2S
 - H_2SO_3
 - H_2SO_4
 - H_2SO_5
 - H_2SO_6

SECTION 8.2 THE NATURE OF COVALENT BONDING

- Draw the Lewis structure for hydrogen fluoride. $\text{H}-\text{F}$
- Draw the Lewis structure for hydrogen chloride. $\text{H}-\text{Cl}$
- Draw the Lewis structure for hydrogen bromine. $\text{H}-\text{Br}$
- Draw the Lewis structure for hydrogen iodine. $\text{H}-\text{I}$
- Draw the Lewis structure for hydrogen cyanide. $\text{H}-\text{C}\equiv\text{N}$
- Draw the Lewis structure for hydrogen cyanide. $\text{H}-\text{C}\equiv\text{N}$



SECTION 8.3 BONDING IN THE CARBON

- Write the Lewis structure for ethane. C_2H_6
- Write the Lewis structure for ethane. C_2H_6

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Liyong Dong



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