

# Chemical Equations and Reactions

QUIZ  
CH #8  
SECTION #1

## Section Quiz: Describing Chemical Reactions

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- B 1. Which of the following is true regarding the law of conservation of mass?
- It dictates that the number of molecules on each side of a chemical equation must be the same.
  - It dictates that the number of atoms of each element must be the same on both sides of a chemical equation.
  - It states that the mass of the reactants must remain constant in order for a chemical reaction to proceed.
  - It does not apply to chemical reactions.
- D 2. Which of the following may indicate that a chemical reaction has occurred?
- release of energy as light
  - a color change
  - gas bubble formation
  - All of the above
- D 3. Which of the following is a product in the reaction described by the word equation below?
- $$\text{iron} + \text{copper(II) sulfate} \rightarrow \text{iron(II) sulfate} + \text{copper}$$
- iron(II) sulfate
  - copper
  - iron
  - Both (a) and (b)
- C 4. What does  $\text{CH}_4$  represent in the balanced chemical equation shown below?
- $$\text{C} + 2\text{H}_2 \rightarrow \text{CH}_4$$
- one atom of  $\text{CH}_4$
  - five atoms of  $\text{CH}_4$
  - one molecule of  $\text{CH}_4$
  - 1 g of  $\text{CH}_4$
- C 5. Which of the following is not true regarding a properly written chemical equation?
- It must be balanced.
  - It must contain proper chemical formulas.
  - It must show the formation of a precipitate or a gas.
  - It should represent known data.

# Chapter 8 Chemical Equations And Reactions Test Answer Key

**Donada Peters, Edward I. Peters**



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## **Table of Contents Chapter 8 Chemical Equations And Reactions Test Answer Key**

1. Understanding the eBook Chapter 8 Chemical Equations And Reactions Test Answer Key
  - The Rise of Digital Reading Chapter 8 Chemical Equations And Reactions Test Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 8 Chemical Equations And Reactions Test Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 8 Chemical Equations And Reactions Test Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 8 Chemical Equations And Reactions Test Answer Key
  - Personalized Recommendations
  - Chapter 8 Chemical Equations And Reactions Test Answer Key User Reviews and Ratings
  - Chapter 8 Chemical Equations And Reactions Test Answer Key and Bestseller Lists
5. Accessing Chapter 8 Chemical Equations And Reactions Test Answer Key Free and Paid eBooks
  - Chapter 8 Chemical Equations And Reactions Test Answer Key Public Domain eBooks
  - Chapter 8 Chemical Equations And Reactions Test Answer Key eBook Subscription Services
  - Chapter 8 Chemical Equations And Reactions Test Answer Key Budget-Friendly Options
6. Navigating Chapter 8 Chemical Equations And Reactions Test Answer Key eBook Formats

- ePub, PDF, MOBI, and More
  - Chapter 8 Chemical Equations And Reactions Test Answer Key Compatibility with Devices
  - Chapter 8 Chemical Equations And Reactions Test Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Chapter 8 Chemical Equations And Reactions Test Answer Key
    - Highlighting and Note-Taking Chapter 8 Chemical Equations And Reactions Test Answer Key
    - Interactive Elements Chapter 8 Chemical Equations And Reactions Test Answer Key
  8. Staying Engaged with Chapter 8 Chemical Equations And Reactions Test Answer Key
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Chapter 8 Chemical Equations And Reactions Test Answer Key
  9. Balancing eBooks and Physical Books Chapter 8 Chemical Equations And Reactions Test Answer Key
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Chapter 8 Chemical Equations And Reactions Test Answer Key
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Chapter 8 Chemical Equations And Reactions Test Answer Key
    - Setting Reading Goals Chapter 8 Chemical Equations And Reactions Test Answer Key
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Chapter 8 Chemical Equations And Reactions Test Answer Key
    - Fact-Checking eBook Content of Chapter 8 Chemical Equations And Reactions Test Answer Key
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
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