

# Calculations need to be based on the limiting reactant.

- Example 2: What is the maximum mass of sulfur dioxide that can be produced by the reaction of 95.6 g carbon disulfide with 100. g oxygen?



$$\frac{95.6 \text{ g } CS_2}{76.2 \text{ g } CS_2} \times \frac{1 \text{ mol } CS_2}{1 \text{ mol } CS_2} \times \frac{2 \text{ mol } SO_2}{1 \text{ mol } CS_2} \times \frac{64.1 \text{ g } SO_2}{1 \text{ mol } SO_2} = 161 \text{ g } SO_2$$

$$\frac{100. \text{ g } O_2}{32.0 \text{ g } O_2} \times \frac{1 \text{ mol } O_2}{3 \text{ mol } O_2} \times \frac{2 \text{ mol } SO_2}{3 \text{ mol } O_2} \times \frac{64.1 \text{ g } SO_2}{1 \text{ mol } SO_2} = 134 \text{ g } SO_2$$

Which reactant is **LIMITING**?

# Limiting Reactant Problems And Solutions

**Huangqi Zhang**



## **Limiting Reactant Problems And Solutions:**

**Multiple Solution Methods for Teaching Science in the Classroom** Stephen DeMeo, 2008 For the first time in science education the subject of multiple solution methods is explored in book form While a multiple method teaching approach is utilized extensively in math education there are very few journal articles and no texts written on this topic in science Teaching multiple methods to science students in order to solve quantitative word problems is important for two reasons First it challenges the practice by teachers that one specific method should be used when solving problems Secondly it calls into question the belief that multiple methods would confuse students and retard their learning Using a case study approach and informed by research conducted by the author this book claims that providing students with a choice of methods as well as requiring additional methods as a way to validate results can be beneficial to student learning A close reading of the literature reveals that time spent on elucidating concepts rather than on algorithmic methodologies is a critical issue when trying to have students solve problems with understanding It is argued that conceptual understanding can be enhanced through the use of multiple methods in an environment where students can compare evaluate and verbally discuss competing methodologies through the facilitation of the instructor This book focuses on two very useful methods proportional reasoning PR and dimensional analysis DA These two methods are important because they can be used to solve a large number of problems in all of the four academic sciences biology chemistry physics and earth science This book concludes with a plan to integrate DA and PR into the academic science curriculum starting in late elementary school through to the introductory college level A challenge is presented to teachers as well as to textbook writers who rely on the single method paradigm to consider an alternative way to teach scientific problem solving

**The Practice of Chemistry Study Guide & Solutions Manual** Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises *Ebook: Chemistry: The Molecular Nature of Matter and Change* Silberberg, 2015-01-16 Ebook Chemistry The Molecular Nature of Matter and Change

**Chemistry, Student Study Guide** James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

**Excel Preliminary Chemistry C. M. Roebuck, 2003** **Problems and Solutions for General Chemistry and College Chemistry, Sixth Editions by Nebergall, Holtzclaw, and**

**Robinson** Frederick K. Ault,1980 *Student Solutions Manual* Darrell Ebbing,2005-03-17 Provides solutions to odd numbered Practice Problems General Problems and Cumulative Skills Problems plus answers to Review Questions

**Solutions Manual to Accompany General Chemistry, Fourth Edition, and General Chemistry with Qualitative Analysis, Fourth Edition, Whitten, Gailey, Davis** Yi-Noo Tang,Wendy L. Keeney-Kennicutt,1992 **Chemistry** Neil D. Jespersen,Alison Hyslop,2021-11-02 Chemistry The Molecular Nature of Matter 8th Edition continues to focus on the intimate relationship that exists between structure at the atomic molecular level and the observable macroscopic properties of matter Key revisions in this edition focus on three areas The deliberate inclusion of more updated real world examples that relate common real world student experiences to the science of chemistry Simultaneously examples and questions have been updated to align them with career concepts relevant to the environmental engineering biological pharmaceutical and medical sciences Providing students with transferable skills with a focus on integrating metacognition and three dimensional learning into the text When students know what they know they are better able to learn and incorporate the material Providing a total solution through New WileyPLUS by fully integrating the enhanced etext with online assessment answer specific responses and additional practice resources The 8th edition continues to emphasize the importance of applying concepts to problem solving to achieve high level learning and increase retention of chemistry knowledge Problems are arranged in an intuitive confidence building order *5 Steps to a 5 on the AP: Chemistry*, **Study and Problem Solving Guide to Accompany Principles of Modern Chemistry, Oxtoby/Nachtrieb** Wade A. Freeman,1986 **Solutions Manual - Chemistry** Charles E. Mortimer,1986 *Student Solutions Manual for Whitten, Davis, Peck, and Stanley's General Chemistry, 7th Ed* Wendy L. Keeney-Kennicutt,2004 Contains answers and solutions to all even numbered end of chapter exercises Solutions are divided by section for easy reference by students **Partial Solutions Manual** Darrell D. Ebbing,David Bookin,1998 *Student Solutions Manual for Chang's Chemistry* Philip C. Keller,Raymond Chang,Jill L. Keller,1988 **Quantitative Chemical Analysis** Daniel C. Harris,Chuck Lucy,2015-05-29 The gold standard in analytical chemistry Dan Harris Quantitative Chemical Analysis provides a sound physical understanding of the principles of analytical chemistry and their applications in the disciplines **CHEMISTRY SILBERBERG**,2003 *Chemistry* Martin Stuart Silberberg,1996 This general chemistry text centres on the theme that observable change in chemical systems is the result of molecular change The aims of this edition are to enable students to perceive matter and change at the molecular level and to help build student confidence in their ability to solve chemical problems as they discover the relevance of chemistry to their lives *Student Solutions Manual: Ssm Chemistry* Deborah Wiegand,2003 This manual contains complete worked out solutions to all follow up problems and about half of all the chapter problems Each chapter of solutions opens with a summary of the text chapter content and a list of key equations needed to solve the problems **Student Solutions Manual for Introduction to General, Organic, and Biochemistry** Morris Hein,2001-07-21 Experienced authors and educators Hein Pattison Best and

Arena recognize that both science and mathematics can be daunting subjects They skillfully anticipate areas of difficulty and pace the text accordingly Particular emphasis is placed on the understanding of how chemical principles relate to the readers lives and future careers

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