

# CHIMICA GENERALE

Ogni atomo/elemento è caratterizzato da una simbolo

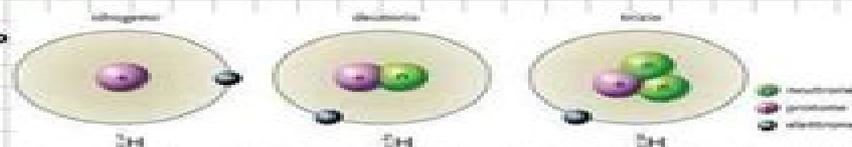


- il numero di massa è dato dalla somma di protoni + neutroni (nucleoni)
- il numero atomico corrisponde al numero dei protoni (uguale al numero di e<sup>-</sup>)

un nucleo con particolari valori di Z e A → **NUCLIDE**

**GLI ISOTOP:**

Atomi di uno stesso elemento che possono differire per il numero di neutroni. I diversi isotopi di un elemento hanno uguali proprietà chimiche, ma diverse proprietà fisiche.



→ **CATIONI:** quando ad un atomo vengono sottratti uno o più elettroni, si formano ioni carichi positivamente.

→ **ANIONI:** quando l'atomo assume uno o più elettroni si formano ioni carichi negativamente.

- contiene 7 periodi (orizzontali)
- 18 gruppi (verticali) → nei gruppi gli elementi hanno caratteristiche chimiche simili

**TAVOLA DEGLI ELEMENTI:**

essi sono organizzati per massa atomica crescente.

gruppi	1A	2A	3A	4A	5A	6A	7A	8A	9A	10A	11A	12A	13A	14A	15A	16A	17A	18A
1	H	He																
2	Li	Be																
3	Na	Mg	Al	Si	P	S	Cl	Ar										
4	K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr
5	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe
6	Cs	Ba	La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn
7	Fr	Ra																

■ orbitale s  
■ orbitale p  
■ orbitale d  
■ orbitale f

- **GRUPPO 1:** metalli alcalini (Li, Na, K, Rb, Cs)
- **GRUPPO 2:** metalli alcalino-terrosi (Be, Mg, Ca, Sr, Ba)
- **GRUPPO 18:** gas nobili, essi si combinano con pochissimi altri elementi. Sono tutti gas incolori e inodori. Sono poco reattivi.
- **GRUPPO 17:** alogeni. Si presentano con molecole biatomiche.

- **I METALLI:** sono solidi, dotati di lucentezza, malleabilità e duttilità; essi sono buoni conduttori di elettricità e calore.

- **I NON METALLI:** sono gas o solidi che non hanno caratteristiche metalliche.

- **SEMIMETALLI O METALLOIDI:** hanno caratteristiche intermedie tra i metalli e i non metalli.

# Appunti Di Chimica Generale Unife

**JL Elias**



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