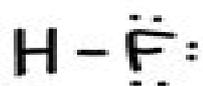
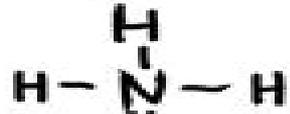
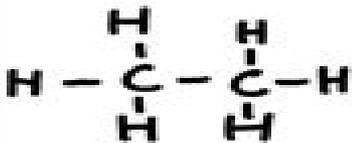
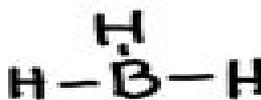
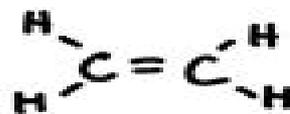
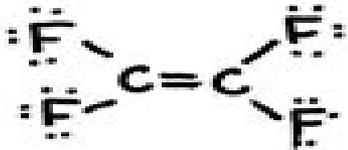


Name Mel

Locker # _____

Bonding & Molecular Structure**LEWIS STRUCTURES**

Indicate the total # of VALENCE electrons for each molecule (VE's). Calculate the number of bonded electrons from the skeleton (BE's). Calculate the Remaining Electrons (RE's = VE's - BE's) and distribute the remaining electrons onto the skeletal structure. Finally, satisfy the octets by creating double and triple bonds. Your final structures should be given below and your scratch paper should clearly show your work.

F_2 VE's = <u>14</u> BE's = <u>2</u> RE's = <u>12</u>  Shape: <u>linear</u>	H_2 VE's = <u>2</u> BE's = <u>2</u> RE's = <u>0</u>  Shape: <u>linear</u>	HF VE's = <u>8</u> BE's = <u>2</u> RE's = <u>6</u>  Shape: <u>linear</u>	NH_3 VE's = <u>8</u> BE's = <u>6</u> RE's = <u>2</u>  Shape: <u>trigonal pyramidal</u>
CH_4 VE's = <u>8</u> BE's = <u>8</u> RE's = <u>0</u>  Shape: <u>tetrahedral</u>	C_2H_6 VE's = <u>14</u> BE's = <u>14</u> RE's = <u>0</u>  Shape: <u>tetrahedral (both)</u>	AlH_3 VE's = <u>6</u> BE's = <u>6</u> RE's = <u>0</u>  Shape: <u>trigonal planar</u>	BH_3 VE's = <u>6</u> BE's = <u>6</u> RE's = <u>0</u>  Shape: <u>trigonal planar</u>
C_2H_4 VE's = <u>12</u> BE's = <u>12</u> RE's = <u>0</u>  Shape around one C: <u>trigonal planar</u>	C_2F_4 VE's = <u>36</u> BE's = <u>12</u> RE's = <u>24</u>  Shape around one C: <u>trigonal planar</u>	CO VE's = <u>10</u> BE's = <u>6</u> RE's = <u>4</u>  Shape: <u>linear</u>	O_2 VE's = <u>12</u> BE's = <u>4</u> RE's = <u>8</u>  Shape: <u>linear</u>

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