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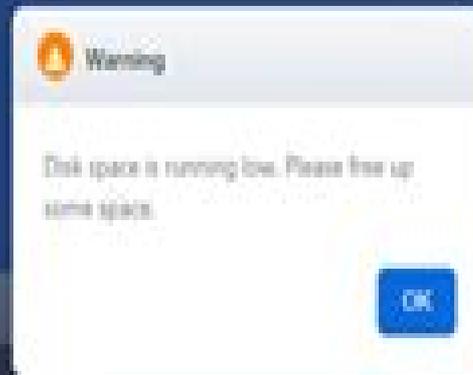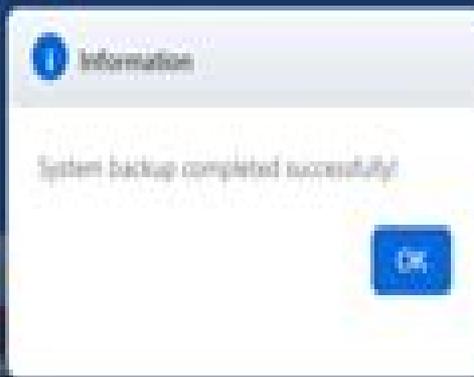
System backup completed successfully!

OK

Warning

Disk space is running low. Please free up some space.

OK

Confirm

Do you want to delete this file?

No   Yes

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web consider using titration to measure the amount of aspirin in a solution using titration it would be difficult to identify the end point because aspirin is a weak acid and reactions may proceed slowly using back titration the end point is more easily recognised in this reaction as it is a reaction between a strong base and a strong acid

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web back titration is used for 1 volatile substances e g nh3 2 insoluble or slightly soluble substances e g caco3 3 substances for which the quantitative reaction proceeds rapidly only in the presence of excess of reagent e g lactic acid aspirin 4 substances which decompose on heating e g

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web titration of aspirin tablets in this lab you will determine the percent purity of two commercially available aspiring tablets using an acid base titration in general an acid and a base react to produce a salt and water by transferring a proton h ha aq naoh aq h2o l naa aq 1 acid base salt

acid base titration of aspirin chemistry stack exchange - Apr 11 2023

web jun 10 2019 1 answer sorted by 3 i am afraid direct acid base titration is not the right way to analyze aspirin exactly because of the you stated hydrolysis the rule number no of any titration is that there should be no side reaction and it

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web determination of an unknown quantity of aspirin using back titration back ground theory many reactions are slow or present unfavourable equilibria for direct titration aspirin is a weak acid therefore it can react with a base such as sodium hydroxide naoh this is an example of an acid base titration see reaction 1

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web topic determination of aspirin in tablets using back titration aim to calculate and compare the active pharmaceutical ingredient api in different commercially available aspirin tablets of the same batch using titration technique standard s indicator s 5 1 8 b1

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web what is back titration it is basically an analytical technique in chemistry which is performed backwardsin the method that is a user needs to find the concentration of a reactant of a given unknown concentration by reacting it with an excess volume of another reactant of a known concentration

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web second you will titrate a sample of your aspirin acetylsalicylic acid with the standardized naoh to determine the moles of acid in a given weight of your product this will allow you to assess its purity assuming the aspirin is not contaminated with other acids the titration allows you to quantitatively determine the purity of your aspirin

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web 4 6 back titration 1 the active ingredient in aspirin tablets is acetylsalicylic acid $c9h8o4$ the acetylsalicylic acid content of an aspirin tablet can be determined using a back titration four aspirin tablets were crushed and added to 25 0 cm3 of 1 00 mol l 1 sodium hydroxide solution

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web pubchemdatabase aspirin cid 2244 here back titration was done to obtain the concentration of aspirin from known aspirin tablets amount of product expected was 0 5g measured yield obtained was 0 335g which shows there was an error in

the determination

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