

# QUALITATIVE ANALYSIS (CATIONS)

| Group | Group reagent  | Cations (Basic radicals) | Composition & colour of precipitate    |
|-------|--|--------------------------|--|
| 1     | dil HCl  | $Pb^{2+}$ (w)            | $PbCl_2$ ; white                       |
|       |  | $Ag^+$                   | $AgCl$ ; white (adv)                   |
|       |  | $Hg_2^{2+}$              | $Hg_2Cl_2$ ; white                     |
| 2(A)  | $H_2S$ in presence of dil HCl (insoluble in YAS)                         | $Pb^{2+}$ (w)            | $PbS$ ; Black                          |
|       |  | $Cu^{2+}$ (w)            | $CuS$ ; Black                          |
|       |  | $Hg^{2+}$                | $HgS$ ; Black (adv)                    |
|       |  | $Bi^{3+}$                | $Bi_2S_3$ ; Black                      |
|       |  | $Co^{2+}$<br>$Cd^{2+}$   | $CoS$ ; yellow                         |
| 2(B)  | $H_2S$ in presence of dil HCl (soluble in YAS)                           | $As^{3+}, Sb^{3+}$       | $As_2S_3$ ; yellow, $Sb_2S_3$ ; orange |
|       |  | $Sn^{2+}, Sn^{4+}$       | $SnS$ ; brown, $SnS_2$ ; yellow        |
| 3     | $NH_4OH$ in presence of $NH_4Cl$   | $Fe^{3+}$ (w)            | $Fe(OH)_3$ ; red brown                 |
|       |  | $Al^{3+}$ (w)            | $Al(OH)_3$ ; gelatinous white          |
|       |  | $Cr^{3+}$                | $Cr(OH)_3$ ; green (adv)               |
| 4     | $H_2S$ in presence of $NH_4OH$ & $NH_4Cl$                                | $Zn^{2+}$ (w)            | $ZnS$ ; white                          |
|       |  | $Ni^{2+}$ (w)            | $NiS$ ; black                          |
| 5     | $(NH_4)_2CO_3$ in presence of $NH_4OH$<br><small>Mns (Bubb/line)</small> | $Mn^{2+}$ (w)            | $MnCO_3$ ; black, $MnS$ ; black        |
|       |  | $Ba^{2+}$ (w)            | $BaCO_3$ ; white                       |
|       |  | $Ca^{2+}$ (w)            | $CaCO_3$ ; white                       |
|       |  | $Sr^{2+}$                | $SrCO_3$ ; white                       |
| 6     | $Na_2HPO_4$ in presence of $NH_4OH$                                      | $Mg^{2+}$ (w)            | $Mg(NH_4)PO_4$ ; white                 |

# Qualitative Analysis Of Cations Lab Report Answers

**George H. Schenk, Darrell D. Ebbing**



## **Qualitative Analysis Of Cations Lab Report Answers:**

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**Qualitative Analysis and Ionic Equilibrium** George H. Schenk, 1990 A supplement for courses with a qualitative analysis component this lab manual contains explanations of the chemistry of metal ions and anions It includes pre lab exercises experiments and lab reports

**Chemistry in the Laboratory** James M. Postma, Julian L. Robert, J. Leland Hollenberg, 2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations

**Abstracts of Papers** American Chemical Society. Meeting, American Chemical Society, 1956

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**Government Reports Index**, 1972 **Processing**, 1979-07 **The Cumulative Book Index**, 1965 A world list of books in the English language **Artificial Intelligence Abstracts**, 1989 *Energy Research Abstracts*, 1980

**Resources in Education**, 1996 *Data Sources*, 2000 **Who's who in American Education** Robert Cecil Cook, 1968

Qualitative Analysis and Ionic Equilibrium George H. Schenk, Darrell D. Ebbing, 1985 A supplement for courses with a qualitative analysis component this lab manual contains explanations of the chemistry of metal ions and anions It includes pre lab exercises experiments and lab reports

**Principles Of Qualitative Inorganic Analysis** Huda Alhasan; Nadiyah, Nadiyah Alahmadi, Huda S Alhasan, 2021-01-26 This book provides notes for basic laboratory experiments in qualitative analysis of cations The book introduces readers to basic methods and laboratory safety Subsequent chapters cover six groups of cations Each chapter explains important details that are required to understand how a particular analytical method works for detecting cations in samples starting from sedimentation and ending with the identification Key Features Simple reader friendly format introductory notes and summary Covers several groups of metals Appendix for handy reference with tables and references This is a useful textbook for early chemistry students and teachers as it equips the readers with sufficient information required to analyze chemical samples and deduce the presence of specific cations as part of laboratory coursework

Nonsulfide Qualitative Analysis of Cations; Separating & Identifying Representative Cations from Group A-E Judith C. Foster, 1995-01-01

**Qualitative Analysis and the Properties of Ions in Aqueous Solution** Emil J. Slowinski, William L. Masterton, 1990 This inexpensive qualitative analysis supplement offers maximum flexibility and

can accompany general chemistry texts Works well with any general chemistry text where the instructor wants more qualitative analysis in conjunction with regular class work

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