



# Geotechnical Engineering Earth Retaining Structures

**RS Peters**



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web 1 molarity  $m = \frac{\text{mass}}{\text{molar mass} \times \text{volume}}$   $0.4500 \text{ L} \times 0.825 \text{ g/L} = 0.37125 \text{ g}$   $0.37125 \text{ g} \times 0.0129 \text{ mol/g} = 0.004789125 \text{ mol}$   $0.004789125 \text{ mol} / 0.00581158 \text{ kg} = 0.825 \text{ mol/kg}$  3 mole fraction  $n_2 = \frac{0.825 \text{ mol}}{0.825 \text{ mol} + 0.00581158 \text{ mol}} = 0.121$   $4 \times 0.825 \text{ g/L} = 3.3 \text{ g/L}$   $3.3 \text{ g/L} \times 18.015 \text{ g/mol} = 59.45 \text{ g/mol}$

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web apr 7 2014 molarity moles of solute liters of solution  $8.4 \times 10^{-2} \text{ mol} / 0.025 \text{ L} = 3.36 \text{ M}$  a first convert 250 ml to liters  $250 / 1000 = 0.25$  then calculate molarity 5 moles / 0.25 liters = 20 M 3 c a solution with molarity 2 requires 2 moles of  $\text{NaOH}$  per liter so  $4 \times 2 = 8 \text{ moles}$  a solution of molarity 1.5 M requires 1.5 mol of  $\text{NaOH}$  to every litre of solvent

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web answer the correct option is c explanation the molarity of a mixture  $M_{\text{mix}}$  can be calculated using the following formula

m mix m 1 v 1 m 2 v 2 v 1 v 2 1 5 x 480 1 2 x 520 480 520 720 624 1000 1 344 m q5 with increase in temperature which of these changes a molality b fraction of solute present in water

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