

# Chapter - 1

## Chemical Reactions and Equations

Chemical Reaction: • Chemical reaction is a process in which new substances with different properties are formed.

- Atoms of one element don't change into another element.
- Only rearrangement of atoms takes place in a chemical reaction.

Examples 1. Hydrogen + Oxygen  $\rightarrow$  water

2. Iron + Oxygen  $\rightarrow$  rust

3. Potassium and chlorine gas  $\rightarrow$  chloride

Physical Change: • No new substance is formed.

- It is a temporary change.
- No change in the composition of the substance takes place.
- It is easily reversible.
- Heat or light may or may not be given out or absorbed.

Chemical Change: • A new substance is formed.

- It is a permanent change.
- The composition of the new substance change greatly.
- The change may be reversible or irreversible.
- Heat or light or both are generally given out or absorbed.

## Law of Conservation of mass

Matter can neither be created nor destroyed in a chemical reaction. This means that total mass of all reactants must

# Chemistry Chapter 1 Notes

**L Darling-Hammond**



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