

Chapter Test II

A. Matching

Match each item in Column I with the correct description in Column II. Write the letter of the correct description.

Column I

- _____ 1. An oxidation-reduction reaction in which an element is oxidized and another element is reduced.
- _____ 2. An equation of a redox reaction that is balanced for both mass and charge.
- _____ 3. The reduction of an oxidizing agent.
- _____ 4. The ratio of the amount of substance that is oxidized to the amount that is reduced.
- _____ 5. An equation that is balanced for both mass and charge.
- _____ 6. An equation that is balanced for mass but not for charge.

Column II

- a. oxidation
- b. net ionic equation
- c. reduction
- d. half-reaction
- e. balanced equation
- f. oxidation-reduction reaction

B. Multiple Choice

Choose the best answer for each question.

- _____ 1. Which of the following is a balanced equation?

<ol style="list-style-type: none"> a. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ b. $2\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$ c. $\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ d. $2\text{H}_2 + 2\text{O}_2 \rightarrow 2\text{H}_2\text{O}$ 	<ol style="list-style-type: none"> e. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}_2$ f. $2\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$ g. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$ h. $2\text{H}_2 + 2\text{O}_2 \rightarrow 2\text{H}_2\text{O}_2$
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- _____ 2. Which of the following is a balanced equation?

<ol style="list-style-type: none"> a. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ b. $2\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$ c. $\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ d. $2\text{H}_2 + 2\text{O}_2 \rightarrow 2\text{H}_2\text{O}$ 	<ol style="list-style-type: none"> e. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}_2$ f. $2\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$ g. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$ h. $2\text{H}_2 + 2\text{O}_2 \rightarrow 2\text{H}_2\text{O}_2$
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- _____ 3. Which of the following is a balanced equation?

<ol style="list-style-type: none"> a. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ b. $2\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$ c. $\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ d. $2\text{H}_2 + 2\text{O}_2 \rightarrow 2\text{H}_2\text{O}$ 	<ol style="list-style-type: none"> e. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}_2$ f. $2\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$ g. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$ h. $2\text{H}_2 + 2\text{O}_2 \rightarrow 2\text{H}_2\text{O}_2$
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