

## SYLLABUS

**Mole Concept and Stoichiometry**

- (i) Gay-Lussac's Law of Combining Volumes; Avogadro's Law.  
*Idea of mole – a number just as dozen, a gross; Avogadro's Law – statement and explanation; Gay-Lussac's Law of Combining Volumes – statement and explanation; "the mass of 22.4 litres of any gas at S.T.P. is equal to its molar mass" (Questions will not be set on formal proof but may be taught for clear understanding) – simple calculations based on the molar volume.*
- (ii) Refer to the atomicity of hydrogen, oxygen, nitrogen and chlorine (proof not required).  
*The explanation can be given using equations for the formation of HCl, NH<sub>3</sub> and NO.*
- (iii) Relative atomic masses (atomic weight) and relative molecular masses (molecular weights): either H = 1 or <sup>12</sup>C = 12 will be accepted; molecular mass = 2 × vapour density (formal proof not required). Deduction of simple (empirical) and molecular formula from the percentage composition of a compound; the molar volume of a gas at S.T.P.; simple calculations based on chemical equations; both reacting weight and volumes.  
*Idea of relative atomic mass and relative molecular mass – standard H atom, or 1/12th of carbon 12 atom.  
 Relating mole and atomic mass, arriving at gram atomic mass and then gram atom; atomic mass is a number dealing with one atom; gram atomic mass is the mass of one mole of atoms.  
 Relating mole and molecular mass arriving at gram molecular mass and gram molecule – molecular mass is a number dealing with a molecule, gram molecular mass is the mass of one mole of molecules.  
 Molecular mass = 2 × vapour density (questions will not be set on formal proof but may be taught for clear understanding); simple calculations based on the formula.  
 Deduction of simple (empirical) and molecular formula from the percentage composition of a compound.*

## 5A. GAY-LUSSAC'S LAW AND AVOGADRO'S LAW

## 5.1 INTRODUCTION

You have learnt in Class IX that all gases behave similarly under similar conditions of temperature and pressure as expressed by gas laws.

Gas laws are certain rules applicable to a gas in respect of change in temperature ( $T$ ) or pressure ( $P$ ) or volume ( $V$ ). The change in any one of these affects the other two.

**Pressure-Volume Relationship or Boyle's Law**

It states that *the volume of a given mass of dry gas is inversely proportional to its pressure at a constant temperature.*

$$P_1 V_1 = P_2 V_2 = k \text{ at constant temperature}$$

**Temperature-Volume Relationship or Charles's Law :**

It states that *volume of a given mass of a dry gas is directly proportional to its absolute (kelvin) temperature, if the pressure is kept constant.*

OR

*The pressure remaining constant, the volume of a given mass of a dry gas increases or decreases by 1/273 of its volume for each 1°C increase or decrease in temperature respectively.*

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} = k \text{ at constant pressure}$$

**Gas Equation**

On combining both the laws we conclude that

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**James E. Brady, John R. Holum**



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science such as chemical equilibria reaction kinetics ideal and non ideal reactor theory and mass accounting by performing practical real world analyses As they progress through the text readers will have the opportunity to analyze a broad range of environmental processes and systems including water and wastewater treatment surface mining agriculture landfills subsurface saturated and unsaturated porous media aqueous and marine sediments surface waters and atmospheric moisture The text begins with an examination of water core definitions and a review of important chemical principles It then progressively builds upon this base with applications of Henry s law acid base equilibria and reactions in ideal reactors Finally the text addresses reactions in non ideal reactors and advanced applications of acid base equilibria complexation and solubility dissolution equilibria and oxidation reduction equilibria Several tools are provided to fully engage readers in mastering new concepts and then applying them in practice including Detailed examples that demonstrate the application of concepts and principles Problems at the end of each chapter challenging readers to apply their newfound knowledge to analyze environmental processes and systems MathCAD worksheets that provide a powerful platform for constructing process models Environmental Process Analysis serves as a bridge between introductory environmental engineering textbooks and hands on environmental engineering practice By learning how to mathematically and numerically model environmental processes and systems readers will also come to better understand the underlying connections among the various models concepts and systems

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