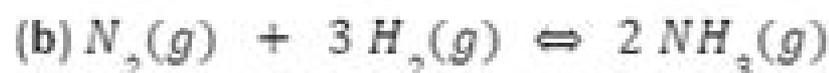


1. Write equilibrium constant expressions, K_p , for the following reactions.



$$\text{Ans. } K_p = \frac{P_{\text{CO}_2} \times P_{\text{H}_2}}{P_{\text{H}_2\text{O}} \times P_{\text{CO}}}$$



$$\text{Ans. } K_p = \frac{P_{\text{NH}_3}^2}{P_{\text{H}_2}^3 \times P_{\text{N}_2}}$$

2. A mixture containing equal concentrations of methane and steam is passed over a nickel catalyst at 1000 K. The emerging gas has the composition $[\text{CO}] = 0.1027 \text{ mol dm}^{-3}$, $[\text{H}_2] = 0.3080 \text{ mol dm}^{-3}$, and $[\text{CH}_4] = [\text{H}_2\text{O}] = 0.8973 \text{ mol dm}^{-3}$. Assuming this mixture is at equilibrium, calculate the equilibrium constant K_c for the reaction

$$\begin{aligned} \text{Ans. } K_c &= \frac{[\text{CO}][\text{H}_2]^3}{[\text{CH}_4][\text{H}_2\text{O}]} = \frac{0.1027 \text{ mol dm}^{-3} \times (0.3080 \text{ mol dm}^{-3})^3}{0.8973 \text{ mol dm}^{-3} \times 0.8973 \text{ mol dm}^{-3}} \\ &= 3.727 \times 10^{-3} \text{ mol}^2 \text{ dm}^{-6} \end{aligned}$$

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Table of Contents Chemical Equilibrium Problems With Solutions

1. Understanding the eBook Chemical Equilibrium Problems With Solutions
 - The Rise of Digital Reading Chemical Equilibrium Problems With Solutions
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Equilibrium Problems With Solutions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Equilibrium Problems With Solutions
 - Personalized Recommendations
 - eBook User Reviews and Ratings
 - eBook Bestseller Lists
5. Accessing Chemical Equilibrium Problems With Solutions Free and Paid eBooks
 - eBook Public Domain eBooks
 - eBook Subscription Services
 - eBook Budget-Friendly Options
6. Navigating eBook Chemical Equilibrium Problems With Solutions eBook Formats
 - ePub, PDF, MOBI, and More
 - eBook Compatibility with Devices
 - eBook Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of eBook Chemical Equilibrium Problems With Solutions
 - Highlighting and Note-Taking eBook Chemical Equilibrium Problems With Solutions
 - Interactive Elements eBook Chemical Equilibrium Problems With Solutions
8. Staying Engaged with eBook Chemical Equilibrium Problems With Solutions

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemical Equilibrium Problems With Solutions
9. Balancing eBooks and Physical Books Chemical Equilibrium Problems With Solutions
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemical Equilibrium Problems With Solutions
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemical Equilibrium Problems With Solutions
- Setting Reading Goals Chemical Equilibrium Problems With Solutions
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemical Equilibrium Problems With Solutions
- Fact-Checking eBook Content of Chemical Equilibrium Problems With Solutions
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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