

4

Atomic Structure
Reviewsheet

A. Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this chapter. Each blank can be completed with a term, short phrase, or number.

Atoms of each element are 1 from the atoms of all other elements. Dalton theorized that atoms are indivisible, but the discovery of 2 particles changed this theory. We now know that atoms are made up of electrons, which have a 3 charge; 4, which have a positive charge; and 5, which are neutral. The latter two particles are found in the 6 of the atom.

It was 7 who discovered the nucleus of the atom. The nucleus has a 8 charge and it occupies a very small volume of the atom. In contrast, the negatively charged 9 occupy most of the volume of the atom.

The number of 10 in the nucleus of the atom is the atomic 11 of that element. Because atoms are electrically neutral, the number of protons and 12 in an atom are equal. The sum of the 13 and neutrons is the mass number. Atoms of the same element are identical in most respects, but they can differ in the number of 14 in the nucleus. Atoms that have the same number of protons but different mass numbers are called 15.

The 16 of an element is the weighted average of the masses of the isotopes of that element. Two isotopes of sulfur are ${}^{32}_{16}\text{S}$ and ${}^{34}_{16}\text{S}$. An atom of the sulfur-32 isotope contains 17 protons and 18 neutrons. The sulfur-34 isotope has 19 protons and 20 neutrons.

1. different 41
2. sub atomic 42
3. negative 42
4. protons 42
5. neutrons 42
6. nucleus 43
7. Rutherford 43
8. positive 43
9. electrons 43
10. protons (elec) 44
11. number 44
12. electrons 44
13. protons (elec) 45
14. neutrons 45
15. isotopes 45
16. atomic mass 47
17. 16 45
18. 16 46
19. 16 46
20. 18 45

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Lingjun Ying



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Table of Contents Chapter 4 Atomic Structure Answers Pdf

1. Understanding the eBook Chapter 4 Atomic Structure Answers Pdf
 - The Rise of Digital Reading Chapter 4 Atomic Structure Answers Pdf
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 4 Atomic Structure Answers Pdf
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 4 Atomic Structure Answers Pdf
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 4 Atomic Structure Answers Pdf
 - Personalized Recommendations
 - Chapter 4 Atomic Structure Answers Pdf User Reviews and Ratings

- Chapter 4 Atomic Structure Answers Pdf and Bestseller Lists
- 5. Accessing Chapter 4 Atomic Structure Answers Pdf Free and Paid eBooks
 - Chapter 4 Atomic Structure Answers Pdf Public Domain eBooks
 - Chapter 4 Atomic Structure Answers Pdf eBook Subscription Services
 - Chapter 4 Atomic Structure Answers Pdf Budget-Friendly Options
- 6. Navigating Chapter 4 Atomic Structure Answers Pdf eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 4 Atomic Structure Answers Pdf Compatibility with Devices
 - Chapter 4 Atomic Structure Answers Pdf Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 4 Atomic Structure Answers Pdf
 - Highlighting and Note-Taking Chapter 4 Atomic Structure Answers Pdf
 - Interactive Elements Chapter 4 Atomic Structure Answers Pdf
- 8. Staying Engaged with Chapter 4 Atomic Structure Answers Pdf
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 4 Atomic Structure Answers Pdf
- 9. Balancing eBooks and Physical Books Chapter 4 Atomic Structure Answers Pdf
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 4 Atomic Structure Answers Pdf
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 4 Atomic Structure Answers Pdf
 - Setting Reading Goals Chapter 4 Atomic Structure Answers Pdf
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 4 Atomic Structure Answers Pdf
 - Fact-Checking eBook Content of Chapter 4 Atomic Structure Answers Pdf
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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