

## 19

## ACIDS, BASES, AND SALTS

**SECTION 19.1 ACID-BASE THEORIES (pages 587–593)**

*This section compares and contrasts acids and bases as defined by the theories of Arrhenius, Brønsted-Lowry, and Lewis. It also identifies conjugate acid-base pairs in acid-base reactions.*

**► Properties of Acids and Bases (pages 587–588)**

- Circle the letters of all the terms that complete the sentence correctly.  
The properties of acids include \_\_\_\_\_ .
  - reacting with metals to produce oxygen
  - giving foods a sour taste
  - forming solutions that conduct electricity
  - causing indicators to change color
- Bases are compounds that react with acids to form \_\_\_\_\_  
and a(n) \_\_\_\_\_ .
- Circle the letters of all the terms that complete the sentence correctly.  
The properties of bases include \_\_\_\_\_ .
  - tasting bitter
  - feeling slippery
  - changing the color of an indicator
  - always acting as a strong electrolyte

**► Arrhenius Acids and Bases (pages 588–590)**

- Match the number of ionizable hydrogens with the type of acid.
 

|             |               |
|-------------|---------------|
| _____ one   | a. diprotic   |
| _____ two   | b. triprotic  |
| _____ three | c. monoprotic |
- Is the following sentence true or false? Only the hydrogens in weak polar bonds are ionizable. \_\_\_\_\_
- Hydrogen is joined to a very \_\_\_\_\_ element in a very polar bond.
- Alkali metals react with water to produce \_\_\_\_\_ solutions.

# Chapter 20 Acids And Bases Answer Key

**Wenbin Ji**



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