

CHAPTER 13

Nuclei

In the last chapter, we learnt that atomic nucleus was discovered in the year 1911 by Rutherford and his associates from the study of large angle scattering of alpha particles from the thin metal foil. In this chapter, we shall study the constituents of the nucleus and how they are held together. We shall also study different properties of nuclei such as size, mass, density and stability of nuclei and associated phenomena such as radioactivity, nuclear fission and nuclear fusion.

NUCLEUS

As per the Rutherford's experiment, it was concluded that the entire positive charge and mass of an atom was present at its very tiny centre known as nucleus.

Composition of nucleus

Nucleus is composed of two types of particles

- Protons** Positively charged particles
- Neutrons** Neutral particles

Mass of neutron is slightly greater than the mass of the proton.

$$\text{Mass of proton, } m_p = 1.6726231 \times 10^{-27} \text{ kg}$$

$$\text{Mass of neutron, } m_n = 1.6749286 \times 10^{-27} \text{ kg}$$

$$m_p \approx 1.8-2.0 m_e, \text{ where } m_e = \text{mass of electron.}$$

Atomic number

Atomic number of an element is the number of proton present inside the nucleus of a neutral atom. It is also equal to the number of electrons revolving in various orbits around the nucleus of an uncharged atom. It is represented by Z .

$$\text{Atomic number (Z) = Number of proton = Number of electron}$$

Mass number

The total number of neutrons and protons, i.e. number of nucleons present inside the nucleus of an element is called its mass number A . Number of neutrons is represented by N .

$$\text{So, } \quad \text{Mass number (A) = Z + N}$$

A chemical element is represented by ${}_Z$ (chemical symbol) ^{A} or ${}_Z X^A$.

Total charge of the nucleus is equal to the total charge on all protons Z in it, i.e. Ze , where $e = 1.6 \times 10^{-19} \text{ C}$.

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Atomic Nuclei Section 1 Answers

RC Schank



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Essential Cell Biology Bruce Alberts, Dennis Bray, Karen Hopkin, Alexander D Johnson, Julian Lewis, Martin Raff, Keith Roberts, Peter

Walter,2015-01-01 Essential Cell Biology provides a readily accessible introduction to the central concepts of cell biology and its lively clear writing and exceptional illustrations make it the ideal textbook for a first course in both cell and molecular biology The text and figures are easy to follow accurate clear and engaging for the introductory student Molecular detail has been kept to a minimum in order to provide the reader with a cohesive conceptual framework for the basic science that underlies our current understanding of all of biology including the biomedical sciences The Fourth Edition has been thoroughly revised and covers the latest developments in this fast moving field yet retains the academic level and length of the previous edition The book is accompanied by a rich package of online student and instructor resources including over 130 narrated movies an expanded and updated Question Bank Essential Cell Biology Fourth Edition is additionally supported by the Garland Science Learning System This homework platform is designed to evaluate and improve student performance and allows instructors to select assignments on specific topics and review the performance of the entire class as well as individual students via the instructor dashboard Students receive immediate feedback on their mastery of the topics and will be better prepared for lectures and classroom discussions The user friendly system provides a convenient way to engage students while assessing progress Performance data can be used to tailor classroom discussion activities and lectures to address students needs precisely and efficiently For more information and sample material visit <http://garlandscience.rocketmix.com>

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Lewars,2010-11-10 This corrected second edition contains new material which includes solvent effects the treatment of

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