



# Sample Detective Case File

**Lingjun Ying**



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web donor whereas a base generically represented by  $b$  is a proton acceptor when acids and bases combine they swap a proton  $ha + b \rightarrow hb$  strong acids e.g. sulfuric acid have a very pronounced tendency to transfer a proton to a base whereas weak acids

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web understand that parts of an acid base titration be able to determine the  $K_a$  or  $K_b$  from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate  $K_{a1}$  and  $K_{a2}$  for a polyprotic acid by the end of this lab students should be able to

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web 1 approximately 3 g of hydrated oxalic acid  $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$  was weighted to the nearest 0.01 g in a 50 ml beaker 2 about 30 ml of distilled water was added to the beaker to dissolve the oxalic acid 3 the solution was transferred into 250 ml volumetric flask the beaker was rinsed and its contents were poured into the flask

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web this problem has been solved you'll get a detailed solution from a subject matter expert that helps you learn core concepts see answer question part b acid base titration lab report i determining the concentration of sulfuric acid solution  $\text{H}_2\text{SO}_4$  by titrating it with  $\text{NaOH}$  solution

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web solution a titrant volume 0 ml the solution pH is due to the acid ionization of  $\text{HCl}$  because this is a strong acid the ionization is complete and the hydronium ion molarity is 0.100 M the pH of the solution is then  $\text{pH} = \log 0.100 = 1.000$  pH log 0.100 = 1.000 b titrant volume 12.50 ml

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web one type of titration uses a neutralization reaction in which an acid and a base react to produce a salt and water in equation 1 the acid is  $\text{HCl}$  hydrochloric acid and the base is  $\text{NaOH}$  sodium hydroxide when the acid and base react they form

nacl sodium chloride which is also known as table salt

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