

# Chemistry: Molality

Directions: Using these formulas, solve the following problems.

$$\text{molality} = \text{mol solute} / \text{kg solvent}$$

$$1 \text{ amu} = 1 \text{ gram} / \text{mol}$$

$$\Delta T_{\text{boiling}} = (i \cdot \text{number of ions}) (\text{molality}) (k_b)$$

$$\Delta T_{\text{freezing}} = (\text{number of ions}) (\text{molality}) (k_f)$$

$k_b$  and  $k_f$  are constants for each individual solvent.

Key



Problem 1: What is the molality of a solution containing 1.23 g of  $\text{Hg}(\text{CN})_2$  dissolved in 1.37 g of methanol? Answer: 3.55 m

Problem 2: What is the molality of a solution containing 16.8 g of NaCl dissolved in 356 g of water? Answer: 0.606 m

Problem 3: How many grams of NaOH would be required to produce a 3.57 m solution with 473 g of water? Answer: 67.6 g

Problem 4: How many grams of KCl would you need to dissolve in 114 g of vinegar in order to prepare a 1.96 m solution? Answer: 16.6 g

Problem 5: Calculate the molality of 5.97 g NaCl dissolved in 471 g of water.  
Answer: 0.217 m

Problem 6: Calculate the molality of 8.91 g  $\text{Na}_2\text{CO}_3$  dissolved in 595 g of water.  
Answer: 0.341 m

Problem 7: Calculate the molality of 8.44 g KBr dissolved in 112 g of water.  
Answer: 0.633 m

Problem 8: What are the new boiling and freezing points for a solution with 581 g KBr dissolved in 845 grams of water? ( $k_b$  water =  $0.512^\circ \text{C/m}$  and  $k_f$  water =  $1.86^\circ \text{C/m}$ )  
Answer: new freezing point =  $-13.4^\circ \text{C}$       new boiling point =  $106^\circ \text{C}$

Problem 9: What are the new boiling and freezing points for a solution with 391 g NaOH dissolved in 1029 grams of water? ( $k_b$  water =  $0.512^\circ \text{C/m}$  and  $k_f$  water =  $1.86^\circ \text{C/m}$ )  
Answer: new freezing point =  $-36.3^\circ \text{C}$       new boiling point =  $107^\circ \text{C}$

# Molality Answers Key

**Steven S. Zumdahl**



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## **Table of Contents Molality Answers Key**

1. Understanding the eBook Molality Answers Key
  - The Rise of Digital Reading Molality Answers Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Molality Answers Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Molality Answers Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Molality Answers Key
  - Personalized Recommendations
  - Molality Answers Key User Reviews and Ratings

- Molality Answers Key and Bestseller Lists
- 5. Accessing Molality Answers Key Free and Paid eBooks
  - Molality Answers Key Public Domain eBooks
  - Molality Answers Key eBook Subscription Services
  - Molality Answers Key Budget-Friendly Options
- 6. Navigating Molality Answers Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Molality Answers Key Compatibility with Devices
  - Molality Answers Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Molality Answers Key
  - Highlighting and Note-Taking Molality Answers Key
  - Interactive Elements Molality Answers Key
- 8. Staying Engaged with Molality Answers Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Molality Answers Key
- 9. Balancing eBooks and Physical Books Molality Answers Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Molality Answers Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Molality Answers Key
  - Setting Reading Goals Molality Answers Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Molality Answers Key
  - Fact-Checking eBook Content of Molality Answers Key
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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