

Prestressed Concrete Design

Second edition

M.K. Hurst



Prestressed Concrete Design Second Edition

Michael Seilmaier



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monoxide CO determine the mass of iodine I_2 which could be produced b if in the above situation only 0.160 moles of iodine i_2
was produced i what mass of iodine was produced

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$$4\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Cu} + 2\text{Ag} + 2\text{Cu(NO}_3)_2$$

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calculated mass of the product expected $\times 100$ in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction $Al_2O_3 + 4 Al + 3 O_2 \rightarrow 2 Al_2O_3$ b determine the theoretical yield of Al_2O_3 160 g Al \times 1 mol Al \times 2 mol Al_2O_3 \times 101.964 g Al_2O_3 302.3 g Al_2O_3 26.982 g Al 4 mol Al 1 mol Al_2O_3 c determine the percent yield actual \times 100 260.0 g \times 100 86.01 theoretical 302.3 g

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