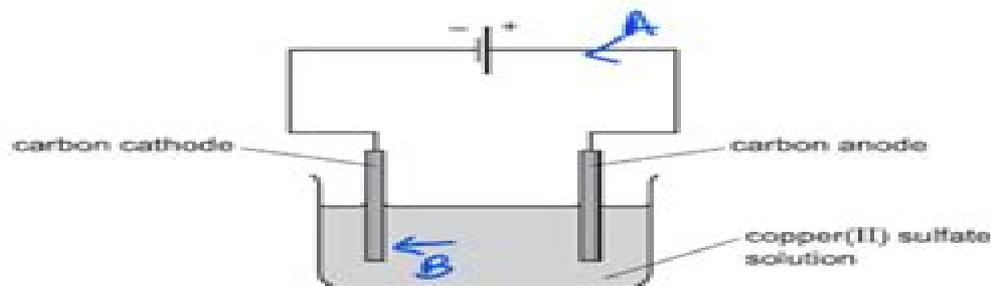


- 4 Copper(II) sulfate solution was electrolysed using the apparatus shown.



- (a) (i) Draw an arrow on the diagram to show the direction of movement of electrons in the wire. Label the arrow **A**.
arrow labelled A on or near wire going in an anti-clockwise direction [1]
- (ii) Draw an arrow on the diagram to show the direction of movement of positive ions in the copper(II) sulfate solution. Label the arrow **B**.
arrow labelled B in electrolyte pointing towards the cathode [1]
- (b) Oxygen was formed at the anode and copper was formed at the cathode.
- (i) The ionic half-equation for the formation of oxygen is shown.



Explain why this reaction is oxidation.

electrons are lost

- [1]
- (ii) Write the ionic half-equation for the formation of copper at the cathode.
M1 Cu²⁺ ions on left
M2 rest of equation correct and correctly balanced
(Cu²⁺ + 2e⁻ → Cu) [2]

- (c) The electrolysis was repeated using copper electrodes in place of carbon electrodes.

State and explain what happens to the masses of the anode and the cathode during this electrolysis.

M1 anode mass decreases

M2 copper lost as ions OR copper (atoms) becomes ions OR



M3 cathode mass increases

M4 copper deposited / layer of copper forms / copper collected at cathode



..... [4]

[Total: 9]

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