

## Calculations in IB Chemistry SL and HL

Below is a list of the calculations are expected to be able to do in IB chemistry. This needs committing to memory, and each type of calculation will need practice to master. Good luck!

Calculations shaded in grey are only required at HL

### Quantitative Chemistry

Purpose	The Calculation	Notes
1. Calculating a quantity in moles from a number of particles	$n = \frac{N}{L}$ <p><math>n</math> = the quantity in moles  <math>N</math> = the number of particles  <math>L</math> = Avogadro's constant, <math>6.02 \times 10^{23}</math></p>	<ul style="list-style-type: none"> <li>When calculating numbers of atoms within molecules, multiply the number of particles (<math>N</math>) by the number of atoms in the formula</li> <li>To find out numbers of particles, rearrange to: <math>N = n \cdot L</math></li> </ul>
2. Determine relative molecular or formula mass, $M_r$	$M_r = \sum (\text{number of atoms} \cdot \text{atomic mass})$	<ul style="list-style-type: none"> <li><math>M_r</math> has no unit as it is a relative value</li> <li>To calculate molar mass, <math>M_{mol}</math>, just stick a 'g' for grams on the end</li> </ul>
3. Calculate a quantity in moles, from a mass of a substance	$n = \frac{m}{M_{mol}}$ <p><math>n</math> = the quantity in moles  <math>m</math> = the mass of substance you are given in grams  <math>M_{mol}</math> = the molar mass of the substance</p>	<ul style="list-style-type: none"> <li>To determine the mass of a given number of moles of a substance use:  <math>m = M_{mol} \cdot n</math></li> <li>To determine the molar mass of a given mass of a substance:  <math>M_{mol} = \frac{m}{n}</math></li> </ul>
4. Determine empirical formula from % composition by mass	<ol style="list-style-type: none"> <li>Divide each % by the atomic mass</li> <li>Divide each answer to Step 1 by the smallest answer to Step 1</li> <li>Multiply all answers to Step 1 to remove any obvious fractions                             <ol style="list-style-type: none"> <li>If there is a '.5' multiply everything by 2</li> <li>If there is a '.3' multiply everything by 3 etc</li> </ol> </li> </ol>	
5. Determine molecular formula from empirical formula	$F_{mol} = \frac{M_r}{m(F_e)} \cdot F_e$ <p><math>F_{mol}</math> = molecular formula  <math>M_r</math> = relative molecular mass  <math>F_e</math> = empirical formula  <math>m(F_e)</math> = empirical formula mass</p>	<ul style="list-style-type: none"> <li>This and the previous calculation are often combined together in exam questions</li> </ul>
6. Use mole ratios to determine the number of moles of B that can be made from A	$n(B) = n(A) \cdot \frac{\text{number of B in equation}}{\text{number of A in equation}}$	<ul style="list-style-type: none"> <li>The second term in this equation is the mole ratio</li> <li>You must use a fully balanced equation</li> <li>This is the central step in many stoichiometry calculations</li> </ul>
7. Calculate theoretical yield	<ol style="list-style-type: none"> <li>Use <b>Calculation 6</b> to determine the expected quantity of product in moles.</li> <li>Use a rearranged <b>Calculation 3</b> to determine the expected mass.</li> </ol>	

# Ib Chemistry Sl Study

**L Towne**



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**Ib Chemistry Sl & Hl** Ivan Reshtenkov, 2017-03-11 This IB Chemistry book may be your best bet for a comprehensive and effective review of the SL or HL course material The book has friendly and understandable explanations of complex concepts with 250 practice questions for the test as well as a complete listing of all related terms and their explanation Important equations are listed throughout each content chapter covering what you need to know in order to excel in the SL or HL test Questions with answers include an overview section and an additional in depth section if you need further clarification The user friendly format makes it one of the best IB Chemistry review books available It provides a means for developing study plans that you can customize to fit your needs It isn't too skimpy or too overwhelming with information It also provides a great way for structuring your studying which is helpful if you consider yourself somewhat less than a totally organized student

*Updraft Downdraft* Marilyn Crawford, Eleanor Dougherty, 2003-06-11 Ever wonder why there is such a big gap in the number of students who succeed in comparison to those who do not The duality of educational experiences that permeate our secondary schools leads to one group of students heading into post secondary education while others either receive terminal degrees or drop out This updraft downdraft phenomena occurs because time talent and monies tend to drift into the updraft group's educational arena leaving the downdraft students and teachers with less time to teach fewer qualified teachers and fewer supporters This book provides a set of tools to assist schools and communities in reallocating their resources more effectively Using school artifacts educators can collect the data and information needed to make better decisions and leverage what they have equitably The process of collecting analyzing and acting on the data described promotes more strategic decision making about school resources Samples of data and analyses around master

schedules student schedules curriculum guides and other artifacts inside the school and district are provided Recommended for anyone seeking ways to understand the problems inside the public schools and ways to find solutions

**Relativistic Quantum Chemistry** Markus Reiher, Alexander Wolf, 2015-01-12 Einstein proposed his theory of special relativity in 1905 For a long time it was believed that this theory has no significant impact on chemistry This view changed in the 1970s when it was realized that nonrelativistic Schrodinger quantum mechanics yields results on molecular properties that depart significantly from experimental results Especially when heavy elements are involved these quantitative deviations can be so large that qualitative chemical reasoning and understanding is affected For this to grasp the appropriate many electron theory has rapidly evolved Nowadays relativistic approaches are routinely implemented and applied in standard quantum chemical software packages As it is essential for chemists and physicists to understand relativistic effects in molecules the first edition of *Relativistic Quantum Chemistry The fundamental Theory of Molecular Science* had set out to provide a concise comprehensive and complete presentation of this theory This second edition expands on some of the latest developments in this fascinating field The text retains its clear and consistent style allowing for a readily accessible overview of the complex topic It is also self contained building on the fundamental equations and providing the mathematical background necessary While some parts of the text have been restructured for the sake of clarity a significant amount of new content has also been added This includes for example an in depth discussion of the Brown Ravenhall disease of spin in current density functional theory and of exact two component methods and its local variants A strength of the first edition of this textbook was its list of almost 1000 references to the original research literature which has made it a valuable reference also for experts in the field In the second edition more than 100 additional key references have been added most of them considering the recent developments in the field Thus the book is a must have for everyone entering the field as well as for experienced researchers searching for a consistent review

**A Comprehensive Treatise on Inorganic and Theoretical Chemistry** Joseph William Mellor, 1922

*Molecular Modelling and Synthesis of Nanomaterials* Ihsan Boustani, 2020-07-14 This book presents nanomaterials as predicted by computational modelling and numerical simulation tools and confirmed by modern experimental techniques It begins by summarizing basic theoretical methods then giving both a theoretical and experimental treatment of how alkali metal clusters develop into nanostructures as influenced by the cluster's magic number of atoms The book continues with a discussion of atomic clusters and nanostructures focusing primarily on boron and carbon exploring in detail the one two and three dimensional structures of boron and carbon and describing their myriad potential applications in nanotechnology from nanocoating and nanosensing to nanobatteries with high borophene capacity The broad discussion of computational modelling as well as the specific applications to boron and carbon make this book an essential reference resource for materials scientists in this field of research

**Chemical Age**, 1924 **Bibliographia Hopkinsiensis 1876-1893** Johns Hopkins University, 1893 *Supplement to Mellor's Comprehensive Treatise on Inorganic and Theoretical*

*Chemistry* Joseph William Mellor, 1922    **A Comprehensive Treatise on Inorganic and Theoretical Chemistry: Ra and Ac families, Be, Mg, Zn, Cd, Hg** Joseph William Mellor, 1923    Supplement to Mellor's Comprehensive Treatise on Inorganic and Theoretical Chemistry: suppl. 3. K, Rb, Cs, Fr Joseph William Mellor, 1922    Johns Hopkins University  
Circulars Johns Hopkins University, 1895    *Circulars* Johns Hopkins University, 1889    **The Johns Hopkins University circular**, 1895    *The Johns Hopkins University Circular* Johns Hopkins University, 1886    **Relativistic Theory of Atoms and Molecules II** Pekka Pyykkö, 1993-11-03 Relativistic effects are of major importance for understanding the properties of heavier atoms and molecules This book is still the only comprehensive bibliography on related calculations The material is organized by subject into tables containing a concise characterization Together with Volume I Lecture Notes in Chemistry Vol 41 ISBN 3 540 17167 3 the literature until 1992 is now covered and 6577 references with titles are given in the two books The book will provide a convenient reference for theoretical chemists and atomic and molecular physicists interested in the properties of heavier elements Contents Introduction One particle problems Quantum electrodynamical effects Multielectron atoms methods Multielectron atoms results Symmetry Molecular calculations Solid state theory Relativistic effects and heavy element chemistry Corrections to Volume I Some comments on notations and terminology List of acronyms and symbols Bibliography    *A Comprehensive Treatise on Inorganic and Theoretical Chemistry: B, Al, Ga, In, Tl, Sc, Ce, and Rare Earth Metals, C (Part I)* Joseph William Mellor, 1924    **The Chemical Age**, 1925    **Bulletin of Carson-Newman College** Carson-Newman College, 1928    Journal of General Chemistry of the USSR in English Translation, 1981-09

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