

### Ideal gas law - problems and solutions

1. Ideal gases in a closed container initially have volume  $V$  and temperature  $T$ . The final temperature is  $5/4T$  and the final pressure is  $2P$ . What is the final volume of gases.

Known:

Initial volume ( $V_1$ ) =  $V$

Initial temperature ( $T_1$ ) =  $T$

Final temperature ( $T_2$ ) =  $5/4 T$

Initial pressure ( $P_1$ ) =  $P$

Final pressure ( $P_2$ ) =  $2P$

Wanted: Final volume ( $V_2$ )

Solution:

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

$$\frac{P V}{T} = \frac{(2P) V_2}{5/4 T}$$

$$\frac{V}{1} = \frac{2 V_2}{5/4}$$

$$5/4 V = 2 V_2$$

$$V_2 = \frac{5V}{4} \cdot \frac{2}{1}$$

$$V_2 = \frac{5V}{4} \times \frac{1}{2}$$

$$V_2 = \frac{5V}{8}$$

2. Determine the volume of 2.00 moles of gases (ideal gas) at STP. STP = Standard Temperature and Pressure.

Known:

Moles of gas ( $n$ ) = 2 moles

Standard temperature ( $T$ ) =  $0^\circ\text{C} = 0 + 273 = 273$  Kelvin

Standard pressure ( $P$ ) = 1 atm =  $1.013 \times 10^5$  Pa

Universal gas constant ( $R$ ) = 8.315 Joule/mole.Kelvin

Wanted: Volume of gases ( $V$ )

Solution:

Read more:

<https://physics.gurumuda.net/ideal-gas-law-problems-and-solutions.htm>

# Gas Law Problems With Solutions

**Didier Musso**



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