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Student Exploration: Covalent Bonds

Vocabulary: covalent bond, diatomic molecule, Lewis diagram, molecule, noble gases, nonmetal, octet rule, shell, valence, valence electron

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

1. There are eight markers in a full set, but Flora and Frank each only have seven markers. Flora is missing the red marker, and Frank is missing the blue marker.

What can they do so that each has a full set of markers? They should share Markers to have a full set.

2. Otto and Olivia each have six markers. Otto is missing the purple and green markers, and Olivia is missing the black and brown markers. What can they do so that each has a full set? The should share markers to have a full set.

Gizmo Warm-up

Just like the students described above, **nonmetal** atoms can share electrons. As you will see in the *Covalent Bonds* Gizmo, atoms form bonds in this way.



To begin, check that **Fluorine** is selected from the **Select a substance** menu. Click **Play** (▶) to see the electrons orbiting the nucleus of each atom.

1. The outermost electrons in each atom are called **valence electrons**. How many valence electrons does each fluorine atom have? 7 Valence electrons
2. Click **Pause** (⏸). Drag a valence electron from the left atom to the right atom. Click **Play**. What happens? A shell is added that connects the two atoms.
3. Click **Pause**, drag a valence electron from the right atom to the left, and then click **Play**. What happens now? An electron is added to the shell that is connecting the two atoms.

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