

Chapter 6 Practice Test: Chemical Bonding

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. The charge on an ion is
a. always positive.
b. always negative.
c. either positive or negative.
d. zero.
- _____ 2. According to the octet rule, a calcium atom has a tendency to
a. lose one electron.
b. lose two electrons.
c. gain one electron.
d. gain two electrons.
- _____ 3. If a compound forms by ionic bonding, which is *not* true?
a. A positively charged atom or group of atoms attracts a negatively charged atom or group of atoms.
b. The net charge of the compound is zero.
c. The compound contains just two atoms, each of opposite charge.
d. Several ions group together in a tightly packed structure.
- _____ 4. The only property listed that is *not* characteristic of ionic compounds is
a. high melting point.
b. hardness.
c. lack of crystal structure.
d. good conductor of electricity.
- _____ 5. Which formula listed below represents a polyatomic ion?
a. NH_4^+
b. H_2SO_4
c. NaCl
d. H_2O
- _____ 6. The melting points of ionic compounds are higher than the melting points of molecular compounds because
a. ionic substances tend to vaporize at room temperature.
b. ionic substances are brittle.
c. attractive forces between ions are greater than the attractive forces between molecules.
d. the numbers of positive and negative charges are equal in an ionic compound.

Modern Chemistry Work Answers Chapter 13

**Christine Lefrou, Pierre Fabry, Jean-
Claude Poignet**



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Electrochemistry Christine Lefrou, Pierre Fabry, Jean-Claude Poignet, 2012-05-24 This textbook offers original and new approaches to the teaching of electrochemical concepts principles and applications Throughout the text the authors provide a balanced coverage of the thermodynamic and kinetic processes at the heart of electrochemical systems The first half of the book outlines fundamental concepts appropriate to undergraduate students and the second half gives an in depth account of electrochemical systems suitable for experienced scientists and course lecturers Concepts are clearly explained and mathematical treatments are kept to a minimum or reported in appendices This book features Questions and answers for self assessment Basic and advanced level numerical descriptions Illustrated electrochemistry applications This book is accessible to both novice and experienced electrochemists and supports a deep understanding of the fundamental principles and laws of electrochemistry *Modern Chemistry* Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001 Modern Electrochemistry 2B John O'M. Bockris, Amulya K. N. Reddy, Maria E. Gamboa-Aldeco, 1998 **Modern University Chemistry, Custom Pub** Norbert T. Porile, 1993 *Lab Experiments for Modern Chemistry* Tzimopoulo, 1990 **Problem Solving Guide and Solutions Manual to Accompany Russell, Fundamentals of Genetics, First Edition** Gail Patt, Anita Bennett, 1994 **American Illustrated Magazine**, 1913 **Modern Chemistry** Nicholas D. Tzimopoulos, 1990 *The Journal of Physical Chemistry*, 1908 **The Journal of Physical Chemistry** Wilder Dwight Bancroft, 1908 Includes section New Books **American Magazine**, 1913 **The American Magazine**, 1913 **Scribner's Monthly** Josiah Gilbert Holland, Richard Watson Gilder, 1913 **Modern Biopharmaceuticals, 4 Volume Set** Jörg Knäblein, 2005-10-28 The biopharmaceutical market has come along way since 1982 when the first biopharmaceutical product recombinant human insulin was launched Over 120 such products are currently being marketed around the world including nine blockbuster drugs The global market for biopharmaceuticals which is currently valued at US 41 billion has been growing at an impressive compound annual growth rate of 21% over the previous five years With over one third of all pipe line products in active development are biopharmaceuticals this segment is set to continue outperforming the total pharmaceutical market and could easily reach US 100 billion by the end of this decade *Nanoscience for the Conservation of Works of Art* Piero Baglioni, David Chelazzi, 2015-11-09 Understanding the chemistry behind works of art and heritage materials presents an opportunity to apply scientific techniques to their conservation and restoration Manipulation of materials at the nanoscale affords greater accuracy and minimal disturbance to the original work while efficiently combating the affects of time and environment This book meets the growing demand for an all encompassing handbook to instruct on the use of today s science on mankind s cultural heritage The editors have pioneered modern techniques in art conservation over the last four decades and have brought together expertise from across the globe Each chapter presents the theoretical background to the topic in question followed by practical information on its application and relevant case studies Introductory chapters present the

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