

**Section 6.1: Organizing the Elements**

- Group 1 on the periodic table is called **Alkali Metals**
- Group 2 on the periodic table is called **Alkaline Earth Metals**
- Elements in groups 3-12 are called **Transition Metals**
- Group 17 on the periodic table is called **Halogens**
- Group 18 on the periodic table is called **Noble Gases**
- Each period on the periodic table corresponds to a **principal energy level**
- The 2 series that make up the inner transition metals are **Lanthanides** and **Actinides**
- Complete the following chart:

Element	Symbol	Group #	Period #	Group Name	Metal, Metalloid, Nonmetal
Zirconium	<b>Zr</b>	<b>4</b>	<b>5</b>	<b>Transition Metals</b>	<b>Metal</b>
<b>Radium</b>	Ra	<b>2</b>	<b>7</b>	<b>Alkaline Earth Metals</b>	<b>Metal</b>
<b>Bismuth</b>	<b>Bi</b>	15	6	<b>Nitrogen Group</b>	<b>Metal</b>
Gallium	<b>Ga</b>	<b>13</b>	<b>4</b>	<b>Boron Group</b>	<b>Metal</b>

- What are triads? **Sets of 3 elements with similar properties** Whose idea were they? **Dobereiner**
- The modern periodic table is arranged by increasing **atomic number**. Whose idea was this? **Moseley**
- Mendeleev** developed the first periodic table, which was arranged by increasing **atomic mass**
- Determine which of the following pairs will have similar properties (Yes or no):
  - Ca and Be? **Yes – same group (Group 2 – Alkaline Earth Metals)**
  - Sr and Y? **No – not the same group**
  - C and Ge? **Yes – same group (Group 14 – Carbon Group)**

**Section 6.3: Periodic Trends**

- As you go across a period, the atomic radius (size) **decreases** because **the attraction between protons and electrons increases, which makes the atom smaller (opposite of tug of war)**
- As you go down a group, the atomic radius (size) **increases** because **the number of principal energy levels increases**
- Circle the atom with the largest atomic radius and put a square around the atom with the smallest radius:
  - |           |           |    |           |
|-----------|-----------|----|-----------|
| <b>Li</b> | <b>Ne</b> | F  | B         |
| <b>N</b>  | <b>Bi</b> | S  | P         |
| Na        | <b>Sr</b> | Al | <b>Cl</b> |
- Ionization energy is the energy required to **remove** an electron.

# Modern Chemistry Chapter 10 Section 5 Review Answers

**P. William Davis, Eldra Pearl  
Solomon, Linda R. Berg**



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