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40 g of H_2O molar mass of H_2O is 18 gm mole and molar mass of CH_3OH is 32 gm mole solution moles of CH_3OH 4.532014 mole moles of H_2O 40.1822 moles therefore according to the equation

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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