

Introduction 11

An Electron's Address

Name: _____
 Date: _____
 Class: _____

Introduction: Energy Levels and Sublevels

In addition to the energy level model, Bohr proposed the following model of energy levels. The energy level of the atom can be thought of as the distance between the nucleus and the energy level. The distance between the nucleus and the energy level is the energy level.

Each energy level contains one or more sublevels. Draw the following table:

Table 1

Energy Level	Sublevels in each energy level
1	s
2	s, p
3	s, p, d
4	s, p, d, f

How does this model describe the energy level of the 1st energy level? How does this model describe the energy level of the 2nd energy level?

Critical Thinking Questions

- How many sublevels are in the 1st energy level?
 The sublevels in the 1st energy level are the 1s sublevel. The energy level is the only sublevel in the 1st energy level.
- How many sublevels are in the 2nd energy level?
 The sublevels in the 2nd energy level are the 2s and 2p sublevels. The energy level is the only sublevel in the 2nd energy level.
- How many sublevels are in the 3rd energy level?
 The sublevels in the 3rd energy level are the 3s, 3p, and 3d sublevels. The energy level is the only sublevel in the 3rd energy level.
- How many sublevels are in the 4th energy level?
 The sublevels in the 4th energy level are the 4s, 4p, 4d, and 4f sublevels. The energy level is the only sublevel in the 4th energy level.
- How many sublevels are in the 5th energy level?
 The sublevels in the 5th energy level are the 5s, 5p, 5d, and 5f sublevels. The energy level is the only sublevel in the 5th energy level.

Chemquest 11 Answers

Michael Rajnik



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