

2.1| INTRODUCTION

In 1809 that John Dalton gave his atomic theory known as Dalton's atomic theory. Dalton considered atom to be indivisible, and indestructible.

Towards the end of 19th century, however, it came to known that atom consists of subatomic particles such as electrons, protons and neutrons.

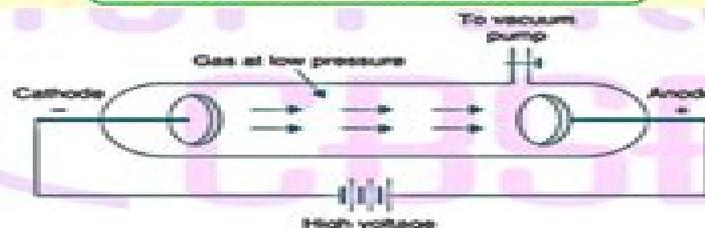
It mean atom was no more indivisible.

SUBATOMIC PARTICLES

- 1) Discovery of Electrons,
- 2) Protons and
- 3) Neutrons

2.2| DISCOVERY OF ELECTRON (CATHODE RAYS): DISCHARGE TUBE EXP.

- (i) A discharge tube is a cylindrical hard glass tube.
- (ii) Fitted with two metallic electrodes.
- (iii) The tube has a side-tube with a stopcock.
- (iv) The discharge tube can be evacuated or filled with any gas to any pressure.
- (v) A high potential difference of 10000V was applied across the electrodes.

WORKING OF EXPERIMENT

- (a) A discharge tube containing air under normal pressure. Electricity does not flow through the gas.
- (b) When the pressure inside tube is reduced to about 10^{-2} atm, the gas starts conducting electricity, and it starts emitting light.

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1. Understanding the eBook Chemistry Notes For Class 11 Cbse Pdf
 - The Rise of Digital Reading Chemistry Notes For Class 11 Cbse Pdf
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Notes For Class 11 Cbse Pdf
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Notes For Class 11 Cbse Pdf
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Notes For Class 11 Cbse Pdf
 - Personalized Recommendations
 - Chemistry Notes For Class 11 Cbse Pdf User Reviews and Ratings
 - Chemistry Notes For Class 11 Cbse Pdf and Bestseller Lists

5. Accessing Chemistry Notes For Class 11 Cbse Pdf Free and Paid eBooks
 - Chemistry Notes For Class 11 Cbse Pdf Public Domain eBooks
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 - Chemistry Notes For Class 11 Cbse Pdf Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
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 - Chemistry Notes For Class 11 Cbse Pdf Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Chemistry Notes For Class 11 Cbse Pdf
 - Highlighting and Note-Taking Chemistry Notes For Class 11 Cbse Pdf
 - Interactive Elements Chemistry Notes For Class 11 Cbse Pdf
8. Staying Engaged with Chemistry Notes For Class 11 Cbse Pdf
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Notes For Class 11 Cbse Pdf
9. Balancing eBooks and Physical Books Chemistry Notes For Class 11 Cbse Pdf
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Notes For Class 11 Cbse Pdf
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Notes For Class 11 Cbse Pdf
 - Setting Reading Goals Chemistry Notes For Class 11 Cbse Pdf
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Notes For Class 11 Cbse Pdf
 - Fact-Checking eBook Content of Chemistry Notes For Class 11 Cbse Pdf
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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