



6: Equilibrium

Chemistry Worksheet

Standard: 11th Science

Subject: Chemistry

Student Name: _____

Q1. Which of the following is not a buffer solution?

1. $\text{CH}_3\text{COOH} + \text{CH}_3\text{COONa}$
2. $\text{H}_2\text{BO}_3 + \text{Na}_2\text{BO}_3$
3. $\text{HClO}_4 + \text{NaClO}_4$
4. $\text{NH}_4\text{OH} + (\text{NH}_4)_2\text{SO}_4$

Q2. 0.10M CH_3COOH is 1.34% ionised, calculate its K_a :

1. 1.8×10^{-1}
2. 1.8×10^{-4}
3. 5×10^{-4}
4. 4×10^{-5}

Q3. Fear or excitement generally causes one to breathe rapidly and it results in the decrease of concentration of CO_2 in blood.

In what way, it will change the pH of blood?

1. pH will increase
2. pH will decrease
3. No change
4. pH will be 7

Q4. Pure ammonia is placed in a vessel at a temperature where its dissociation constant (α) is appreciable. At equilibrium :

1. K_p does not change significantly with pressure.
2. α does not change with pressure.
3. The concentration of NH_3 does not change with pressure.
4. The concentration of hydrogen is less than that of nitrogen.

Q5. Which of the following salts will give highest pH in water?

1. KCl
2. NaCl
3. Na_2CO_3
4. CuSO_4

Q6. "An acid is a substance that is capable of donating a hydrogen ion H^+ and bases are substances capable of accepting a



Chemistry Chemical Equilibrium Test Answers

Alexander Smith, James Kendall



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CliffsNotes AP Chemistry Bobrow Test Preparation Services,2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter s end Two full length practice tests with detailed answer explanations are included in the book

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