

Chemistry Worksheets Class 10 on Chapter 1 Chemical Reactions and Equations- Set 1

Q-1: Write the balanced chemical equations for each of the following reactions.

- Liquid hydrogen peroxide decomposes to form water and oxygen gas.
- Acetylene gas burns in oxygen to form carbon dioxide and water along with evolution of heat.
- When solid mercury(II) oxide is heated, liquid mercury and oxygen gas are produced.
- Marble dissolves in a hydrochloric acid to give calcium chloride, water and carbon dioxide.
- Copper sulphate on treatment with potassium iodide precipitates cuprous iodide, liberates iodine gas and also forms potassium sulphate.

Q-2: Fill in the blanks

- _____ is the magnetic oxide of iron.
- In reactivity series, manganese lies _____ zinc.
- To the reaction between AgNO_3 solution and NaCl solution, the names given to this type of reactions are _____ and _____.
- Proteins decompose to form _____ in our body.
- If copper turnings are added to colourless AgNO_3 solution, it turns _____ after sometime.

Q-3: Which of the following displacement reactions will not take place?

- $\text{Zn(s)} + \text{FeSO}_4(\text{aq}) \rightarrow \text{ZnSO}_4(\text{aq}) + \text{Fe(s)}$
- $\text{Fe(s)} + \text{CuSO}_4(\text{aq}) \rightarrow \text{FeSO}_4(\text{aq}) + \text{Cu(s)}$
- $\text{Cu(s)} + 2\text{AgNO}_3(\text{aq}) \rightarrow \text{Cu(NO}_3)_2(\text{aq}) + 2\text{Ag(s)}$
- $\text{Cu(s)} + \text{FeSO}_4(\text{aq}) \rightarrow \text{CuSO}_4(\text{aq}) + \text{Fe(s)}$

Q-4: Which of the following statements is false?

- Oxidation results in the loss of an electropositive element and the gain of an electronegative element.
- Reduction results in the loss of an electronegative element and the gain of an electropositive element.
- A species that gains electrons is referred to as a reducing agent.
- An oxidising agent is a type of specie that absorbs hydrogen atoms.

Q-5: Why are certain chemicals and medications kept in coloured bottles?

Q-6: Sakshi desired that her home be white washed. She went to the market and bought 20 kg of quick lime, which she dissolved in 60 litres of water. When she added lime to water, she noticed that the water began to boil even though it was not heated. Give an explanation for her observation. Name the product and write the corresponding chemical equation and word equation.

Chemistry Chapter 10 Answers

Thomas Evangelist



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Table of Contents Chemistry Chapter 10 Answers

1. Understanding the eBook Chemistry Chapter 10 Answers
 - The Rise of Digital Reading Chemistry Chapter 10 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 10 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Chapter 10 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 10 Answers
 - Personalized Recommendations
 - Chemistry Chapter 10 Answers User Reviews and Ratings

- Chemistry Chapter 10 Answers and Bestseller Lists
- 5. Accessing Chemistry Chapter 10 Answers Free and Paid eBooks
 - Chemistry Chapter 10 Answers Public Domain eBooks
 - Chemistry Chapter 10 Answers eBook Subscription Services
 - Chemistry Chapter 10 Answers Budget-Friendly Options
- 6. Navigating Chemistry Chapter 10 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Chapter 10 Answers Compatibility with Devices
 - Chemistry Chapter 10 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter 10 Answers
 - Highlighting and Note-Taking Chemistry Chapter 10 Answers
 - Interactive Elements Chemistry Chapter 10 Answers
- 8. Staying Engaged with Chemistry Chapter 10 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 10 Answers
- 9. Balancing eBooks and Physical Books Chemistry Chapter 10 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 10 Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Chapter 10 Answers
 - Setting Reading Goals Chemistry Chapter 10 Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Chapter 10 Answers
 - Fact-Checking eBook Content of Chemistry Chapter 10 Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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