

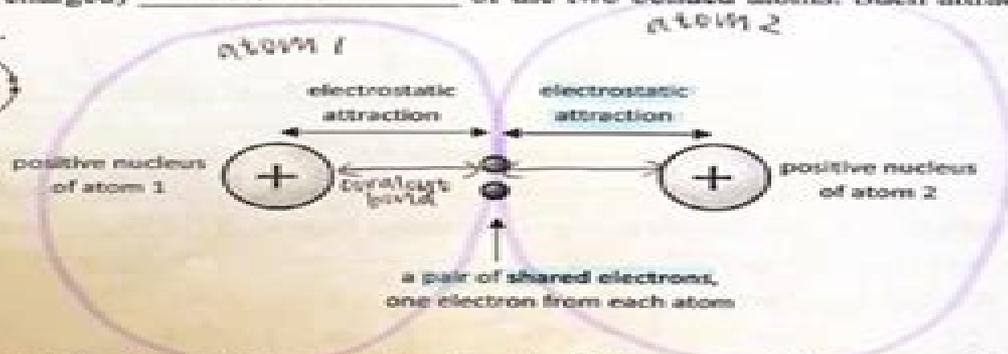
Chapter 8: Covalent Bonding

Students should:

- 1.14 know what is meant by the term molecule
- 1.44 know that a covalent bond is formed between atoms by the sharing of a pair of electrons
- 1.45 understand covalent bonds in terms of electrostatic attractions
- 1.46 understand how to use dot-and-cross diagrams to represent covalent bonds in:
 - diatomic molecules, including hydrogen, oxygen, nitrogen, halogens and hydrogen halides
 - inorganic molecules including water, ammonia and carbon dioxide
 - organic molecules containing up to two carbon atoms, including methane, ethane, ethene and those containing halogen atoms
- 1.47 explain why substances with simple molecular structures are gases or liquids, or solids with low melting and boiling points
the term intermolecular forces of attraction can be used to represent all forces between molecules
- 1.48 explain why the melting and boiling points of substances with simple molecular structures increase, in general, with increasing relative molecular mass
- 1.49 explain why substances with giant covalent structures are solids with high melting and boiling points
- 1.50 explain how the structures of diamond, graphite and C₆₀ fullerene influence their physical properties, including electrical conductivity and hardness
- 1.51 know that covalent compounds do not usually conduct electricity

8.1 Covalent bonding

A covalent bond is resulted when non-metal atoms combine with non-metal atoms by sharing of outermost shell electrons. Strong, directional electrostatic attraction is formed between the (negatively charged) shared electrons and the (positively charged) nuclei of the two bonded atoms. Such attraction is called a



The bonded atoms can then attain an octet or duplet of electrons and become stable (i.e. each atom attains the stable electronic structure of its nearest noble gas).

Chapter 8 Covalent Bonding Assessment Answers

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Chapter 8 Covalent Bonding Assessment Answers:

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General Chemistry Raymond Chang,2003 Publisher Description **Holt Chemistry** Ralph Thomas Myers,2004

Basic Chemistry for Biology Carolyn Chapman,1996 This supplement is for life science majors taking general biology who lack a basic understanding of inorganic chemistry **Medical Pharmacology and Therapeutics E-Book** Derek G. Waller,Andrew W. Hitchings,2021-10-20 *Medical Pharmacology* drug actions side effects and the therapeutic principles of drug use It takes a helpful systems based approach that orders information according to body systems and disease areas rather than by drug class Now in its sixth edition the book has been fully updated to include latest scientific understanding of drug action and administration and current best practice in prescribing medications informed by the latest national guidelines A clinical focus throughout suitable for medical nursing and other healthcare students throughout their training Thorough update of clinically relevant medicines to ensure best practice Information aligned with the British National Formulary BNF NICE guidelines and relevant professional benchmarks Comprehensive drug compendia allow all drugs to be

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