

## Lab 22A

### Investigating Chemical Equilibrium

**Purpose:** To investigate the macroscopic properties of two chemical systems at equilibrium, observe shifts in equilibrium concentrations as a function of concentration and temperature and relate the to systems, which are in equilibrium, their structures through applying Le Chatelier's principle.

#### Proble:

1. Write a balanced equation for the reaction in part 1.  
 $\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$
2. Write a balanced equation for the reaction in part 2.  
 $\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$
3. Write a balanced equation for the reaction in part 3.  
 $\text{Co}^{2+}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \text{CoCl}_4^{2-}(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$
4. In part 1, what is the conjugate acid and base? Write the  $K_{\text{eq}}$  for the reaction. Is the reaction favored to the right or left?  
Conjugate acid:  $\text{HSCN}$  and conjugate base:  $\text{SCN}^{-}$ . The  $K_{\text{eq}}$  is  $K_{\text{eq}} = \frac{[\text{FeSCN}^{2+}]}{[\text{Fe}^{3+}][\text{SCN}^{-}]}$ . The reaction is favored to the right.
5. What is the conjugate acid and base? Write the  $K_{\text{eq}}$  for the reaction. Is the reaction favored to the right or left?  
Conjugate acid:  $\text{HSCN}$  and conjugate base:  $\text{SCN}^{-}$ . The  $K_{\text{eq}}$  is  $K_{\text{eq}} = \frac{[\text{FeSCN}^{2+}]}{[\text{Fe}^{3+}][\text{SCN}^{-}]}$ . The reaction is favored to the right.
6. Write the equation for the reaction between  $\text{Fe}^{3+}$  and  $\text{OH}^{-}$  ions.  
 $\text{Fe}^{3+} + 3\text{OH}^{-} \rightleftharpoons \text{Fe}(\text{OH})_3$
7. Write the equation for the reaction between  $\text{Fe}^{3+}$  and  $\text{OH}^{-}$  ions.  
 $\text{Fe}^{3+} + 3\text{OH}^{-} \rightleftharpoons \text{Fe}(\text{OH})_3$
8. Write the equation for the reaction between  $\text{Fe}^{3+}$  and  $\text{OH}^{-}$  ions.  
 $\text{Fe}^{3+} + 3\text{OH}^{-} \rightleftharpoons \text{Fe}(\text{OH})_3$
9. Write the equation for the reaction between  $\text{Fe}^{3+}$  and  $\text{OH}^{-}$  ions.  
 $\text{Fe}^{3+} + 3\text{OH}^{-} \rightleftharpoons \text{Fe}(\text{OH})_3$

# Investigating Chemical Equilibrium Lab 12a Answer

**ML Yell**



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