

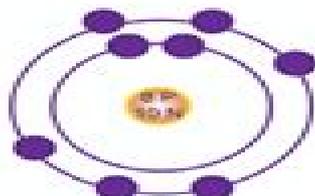
# Average Atomic Mass Lab

M. Lopatka WPHS

## “Beanium”



$^{16}\text{O}$  Isotope



$^{18}\text{O}$  Isotope

**Read** the information below and follow the directions to complete the lab. This is a group assignment, as all labs are, but make sure **all** partners participate equally as lab questions also **appear on tests, quizzes, and the CFE.**

**All** elements on the Periodic Table exist in at least two isotopic forms. Isotopes are atoms with the same atomic number but with different mass numbers due to varying numbers of neutrons. The atomic mass shown on the Periodic Table for each

element is actually an average of all the isotopes of that element, weighted by the percentage of the abundance in which they occur. It should be noted that different isotopes of an element are not found in different places around the universe. In reality, every sample of that element contains all the different isotopes in fixed percentages, no matter where in the universe the sample is found.

### NEWS FLASH!!! A NEW ELEMENT HAS BEEN DISCOVERED.

**Winter Park, FL USA** – Nuclear chemists, performing basic research on food products at a Winter Park Power Plant, have discovered what is believed to be a new element. Mr. Cool Guy, the plant’s owner, says, “We have tentatively named this element Beanium (Bean ee umm).” Mr. Cool Guy’s, assistant Ms. Whiz kid, adds, “We derived this element from the protein nodules found in many types of legumes (bean plants).”

**Further** research of the new element will be conducted in more suitable surroundings, namely laboratories in Wildcat Country, the legendary Winter Park High school. This research work has been contracted to Mr. Lopatka’s amazing chemistry students at WPHS. “Student excitement regarding this discovery is running at an astonishing high rate!” says Lisa Smart Kid, a WPHS student. Many chemistry students have generously volunteered their time and expertise to help with the follow-up experiments involving the new element.

**Dr. Great Chemist** says the first follow-up experiments conducted at Winter Park High School will determine how many isotopes of this element exist. The second experiment will determine the mass of each isotope. The third experiment will determine the percent abundance of each isotope. **And the final calculations will discover the average atomic mass of the new element.**

“**One** unique property of Beanium should make these experiments particularly easy – unlike normal atoms, Beanium atoms are very large,” says Mr. Lopatka. “They can be easily seen, and different isotopes can be sorted by hand.”

**Scientists** are expecting a complete, comprehensive summary of this new element within two days, including diagrams and collected data tables.

# Bean Lab Answers

**B Lingard**



## Bean Lab Answers:

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