

UNIT-I

WATER TREATMENT AND TECHNOLOGY

1.1 INTRODUCTION

Water is essential for the existence of human beings, animals and plants. Though 80% of the earth's surface is occupied by water, less than 1% of the water is available for ready use.

The main sources of water are

- Rain
- rivers and lakes (surface water)
- wells and springs (ground water)
- sea water

Among the above sources of water, rain water is the purest form of water but it is very difficult to collect whereas sea water is the most impure form. Thus, surface and ground water are normally used for industrial and domestic purposes. Such water must be free from undesirable impurities.

Water Treatment

"The process of removing all types of impurities from water and making it fit for domestic or industrial purposes are called Water treatment or Water Technology"

1.2 HARDNESS OF WATER

Definition: If the water produces lather with soap, it is **soft water**. If the water does not produce lather with soap, it is called as **Hard water**. The property is known as hardness. But it will produce a scummy white precipitate. The hardness is due to Ca^{2+} , Mg^{2+} , Cl^- , CO_3^{2-} , HCO_3^- ions and their salts.

Test for hardness: a) Eriochrome Black – T indicator gives wine red colour in hard water.

b) With soap, hard water gives a scummy precipitate.

Types of Hardness:

- a) **Temporary Carbonate hardness:** The hardness due to carbonates and bicarbonates can be removed by simple boiling. So, they are known as temporary or carbonate or alkaline hardness.



(b). **Permanent hardness:** (Non-carbonate): The hardness due to chlorides and sulphates cannot be removed by simple boiling. They need special methods like lime soda process and zeolite process. This hardness is called as "Permanent hardness or Noncarbonate or non alkaline hardness.



Units of hardness: i) ppm ii) mg/L iii) degree Clarkes iv) Degree French

Engineering Chemistry 1 Water Unit Notes

Christian Drosten



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Table of Contents Engineering Chemistry 1 Water Unit Notes

1. Understanding the eBook Engineering Chemistry 1 Water Unit Notes
 - The Rise of Digital Reading Engineering Chemistry 1 Water Unit Notes
 - Advantages of eBooks Over Traditional Books
2. Identifying Engineering Chemistry 1 Water Unit Notes
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Engineering Chemistry 1 Water Unit Notes
 - User-Friendly Interface
4. Exploring eBook Recommendations from Engineering Chemistry 1 Water Unit Notes
 - Personalized Recommendations
 - Engineering Chemistry 1 Water Unit Notes User Reviews and Ratings
 - Engineering Chemistry 1 Water Unit Notes and Bestseller Lists
5. Accessing Engineering Chemistry 1 Water Unit Notes Free and Paid eBooks
 - Engineering Chemistry 1 Water Unit Notes Public Domain eBooks
 - Engineering Chemistry 1 Water Unit Notes eBook Subscription Services
 - Engineering Chemistry 1 Water Unit Notes Budget-Friendly Options
6. Navigating Engineering Chemistry 1 Water Unit Notes eBook Formats

- ePub, PDF, MOBI, and More
 - Engineering Chemistry 1 Water Unit Notes Compatibility with Devices
 - Engineering Chemistry 1 Water Unit Notes Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Engineering Chemistry 1 Water Unit Notes
 - Highlighting and Note-Taking Engineering Chemistry 1 Water Unit Notes
 - Interactive Elements Engineering Chemistry 1 Water Unit Notes
 8. Staying Engaged with Engineering Chemistry 1 Water Unit Notes
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Engineering Chemistry 1 Water Unit Notes
 9. Balancing eBooks and Physical Books Engineering Chemistry 1 Water Unit Notes
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Engineering Chemistry 1 Water Unit Notes
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Engineering Chemistry 1 Water Unit Notes
 - Setting Reading Goals Engineering Chemistry 1 Water Unit Notes
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Engineering Chemistry 1 Water Unit Notes
 - Fact-Checking eBook Content of Engineering Chemistry 1 Water Unit Notes
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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