

Physical Science – 2nd Semester – Final Exam Study Guide – (2012–2013)**1. Physical Science Basics**

- a. What tool(s) would you want to use to find the:
- Mass of an object? **triple beam balance** Basic SI Unit? **kg**
 - Volume of an object? **ruler or water displacement method** Basic SI Unit? **cm³ or mL**
 - Length of an object? **Ruler** Basic SI Unit? **meter**
 - Density of an object? **graduated cylinder** Basic SI Unit? **mL**
- b. What is the formula for calculating density? **mass/volume**
- What is the density of a cube that is 35 g and has a volume of 3 cm³? **35/3cm³**
 - Compare the density of solid H₂O to liquid H₂O, and explain what observable effect this has. **solid H₂O less dense than liquid H₂O**

2. Chemistry Basics

- a. Element (def.) **substance made up of one kind of atom**
- b. Atom (def.) **basic building block of matter**
- c. Compound (def.) **a molecule that contains atoms or molecules of two or more different elements**
- d. Molecule (def.) **a combination of two or more of the same or different atoms that are bonded together**

3. Model of the Atomic Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created **models** to describe them.

Why have atomic theories changed over time? **better technology came around.**

| Scientist | Year | Major Discovery/ Contribution | Name of atomic model |
|-------------------|------|-------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|----------------------|
| John Dalton | 1808 | Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses. | Billiard Ball Model |
| J.J. Thomson | 1897 | Negatively charged electron. | Plum Pudding Model |
| Niels Bohr | 1913 | Electrons move fixed orbits around the atom's nucleus in specific layers, or electron shells . Atoms absorb or give off energy when the electrons move from one shell (energy level) to another. | Bohr Model |
| Ernest Rutherford | 1909 | 1909 (Based on idea that positive reflects positive) atom mostly empty space; electrons orbit randomly around positively charged nucleus. | Planetary Model |

4. Periodic Table

- a. The current periodic table is arranged by increasing **atomic number**.
- b. Reading the Periodic Table
- An element's **properties** can be predicted from its location in the periodic table.
 - The atomic number represents the number of **protons**.
 - The atomic mass represents the number of **protons and neutrons**.

Physical Science Final Exam 3 Study Guide

Christian G. Meyer



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