

# Atomic structure and the periodic table

## Atomic structure



protons  
neutrons  
electrons

atoms have no overall charge

particle	relative mass	relative charge
proton	1	+1
neutron	1	0
electron	very small	-1

## elements

elements - substances made up of atoms with the same atomic number

isotopes of an element - atoms with the same number of protons but different numbers of neutrons

relative atomic mass - the average mass number for an element

## Equations, mixtures and chromatography

### chemical formulas and equations

chemical formula - shows the proportion of atoms of each element in a compound

chemical equation - shows the overall change in a reaction

example:

reactants

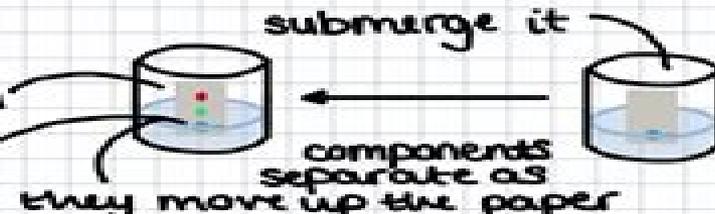
products

word equation: methane + oxygen → carbon dioxide + water

symbol equation:  $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$

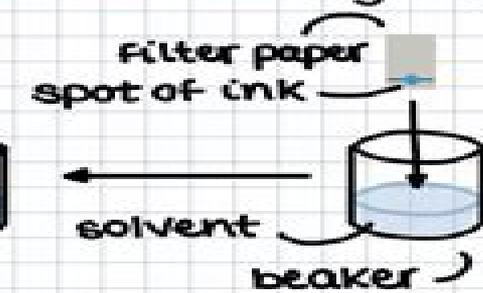
chromatogram

insoluble components stay on the base line



submerge it  
components separate as they move up the paper

### paper chromatography



## nuclear symbols

mass number = total number of protons and neutrons in an atom



atomic number = number of protons in an atom

## compounds

compound - a substance formed from two or more elements chemically bonded together in fixed proportions

molecule - particle containing two or more non-metal atoms bonded covalently. Molecules can be elements or compounds

## mixtures

mixtures - substances made up of different elements or compounds that are not chemically bonded to each other

mixtures can be separated using physical methods

# Chapter 7 Notes Atomic Structure And Periodicity

**Paul B. Kelter**



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## **Table of Contents Chapter 7 Notes Atomic Structure And Periodicity**

1. Understanding the eBook Chapter 7 Notes Atomic Structure And Periodicity
  - The Rise of Digital Reading Chapter 7 Notes Atomic Structure And Periodicity
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 7 Notes Atomic Structure And Periodicity
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 7 Notes Atomic Structure And Periodicity
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 7 Notes Atomic Structure And Periodicity
  - Personalized Recommendations
  - Chapter 7 Notes Atomic Structure And Periodicity User Reviews and Ratings
  - Chapter 7 Notes Atomic Structure And Periodicity and Bestseller Lists
5. Accessing Chapter 7 Notes Atomic Structure And Periodicity Free and Paid eBooks
  - Chapter 7 Notes Atomic Structure And Periodicity Public Domain eBooks
  - Chapter 7 Notes Atomic Structure And Periodicity eBook Subscription Services
  - Chapter 7 Notes Atomic Structure And Periodicity Budget-Friendly Options

6. Navigating Chapter 7 Notes Atomic Structure And Periodicity eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 7 Notes Atomic Structure And Periodicity Compatibility with Devices
  - Chapter 7 Notes Atomic Structure And Periodicity Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 7 Notes Atomic Structure And Periodicity
  - Highlighting and Note-Taking Chapter 7 Notes Atomic Structure And Periodicity
  - Interactive Elements Chapter 7 Notes Atomic Structure And Periodicity
8. Staying Engaged with Chapter 7 Notes Atomic Structure And Periodicity
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 7 Notes Atomic Structure And Periodicity
9. Balancing eBooks and Physical Books Chapter 7 Notes Atomic Structure And Periodicity
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 7 Notes Atomic Structure And Periodicity
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter 7 Notes Atomic Structure And Periodicity
  - Setting Reading Goals Chapter 7 Notes Atomic Structure And Periodicity
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 7 Notes Atomic Structure And Periodicity
  - Fact-Checking eBook Content of Chapter 7 Notes Atomic Structure And Periodicity
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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