

Name _____

Multiple Choice (10 points each)

Identify the letter of the choice that best completes the statement or answers the question.

- E 1. For a certain process at 27°C, $\Delta G = +210.6$ kJ and $\Delta H = -168.2$ kJ. What is the entropy change for this process at this temperature? Express your answer in the form, $\Delta S =$ _____ J/K.
- a. +628.3 J/K
b. -141.3 J/K
c. +141.3 J/K
d. 1.26×10^3 J/K
e. -1.26×10^3 J/K

Answer: $\Delta G = \Delta H - T\Delta S$ $T = 27^\circ\text{C} + 273 = 300\text{K}$

$$\Delta S = \frac{\Delta H - \Delta G}{T} = \frac{(-168.2\text{kJ}) - (+210.6\text{kJ})}{300\text{K}} = -\frac{1.26\text{kJ}}{\text{K}}$$

$$\left(\frac{-1.26\text{kJ}}{\text{K}}\right)\left(\frac{10^3\text{J}}{\text{kJ}}\right) = \boxed{\frac{-1.26 \times 10^3\text{J}}{\text{K}}}$$

- A 2. Which response includes all of the following substances that have $\Delta H_f^\circ = 0$, and no other substances?

- I. HCl(g)
II. Na(s)
III. HCl(aq)
IV. F₂(g)

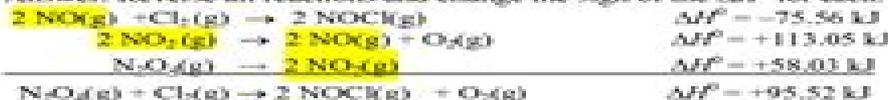
- a. II and IV
b. II
c. I and II
d. I, II, III, and IV
e. I, II, and IV

Answer: All elements in their natural standard state have a $\Delta H_f^\circ = 0$.

- A 3. From the following data at 25°C,

Calculate ΔH° at 25°C for the reaction below.

- a. +95.52 kJ
b. +299 kJ
c. -186.8 kJ
d. -85.52 kJ
e. -246.65 kJ

Answer: Reverse all reactions and change the sign of the ΔH° for each.

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ML Yell



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Commonly Asked Questions in Thermodynamics Marc J. Assael, William A. Wakeham, Anthony R. H. Goodwin, Stefan Will, Michael Stamatoudis, 2011-03-10. Have you ever had a question that keeps persisting and for which you cannot find a clear answer? Is the question seemingly so simple that the problem is glossed over in most resources or skipped entirely? CRC Press Taylor and Francis is pleased to introduce *Commonly Asked Questions in Thermodynamics*, the first in a new series of books that address the questions that frequently arise in today's major scientific and technical disciplines. Designed for a wide audience from students and researchers to practicing professionals in related areas, the books are organized in a user-friendly Question Answer format. Presented questions become increasingly specific throughout the book with clear and concise answers as well as illustrations, diagrams and tables are incorporated wherever helpful. Thermodynamics is a core discipline associated with the theoretical principles and practical applications underlying almost every area of science from nanoscale biochemical engineering to astrophysics. Highlighting chemical thermodynamics in particular, this book is written in an easy to understand style and provides a wealth of fundamental information, simple illustrations and extensive references for further research and collection of specific data. Designed for an audience that ranges from undergraduate students to scientists and engineers at the forefront of research, this indispensable guide presents clear explanations for topics with wide applicability. It reflects the fact that very often the most common questions are also the most profound. *The Thermodynamics of Electrical Phenomena in Metals, and A Condensed Collection of Thermodynamic Formulas* Percy Williams Bridgman, 1961. [Physical Chemistry Student Solutions Manual](#) Charles Trapp, Peter Atkins, Marshall Cady, 2006-08-11. Change 21.

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Problems in Thermodynamics and Statistical Physics Peter T. Landsberg,2014-07-16 Well respected and widely used this volume presents problems and full solutions related to a wide range of topics in thermodynamics statistical physics and statistical mechanics The text is intended for instructors undergraduates and graduate students of mathematics physics chemistry and engineering Twenty eight chapters each prepared by an expert proceed from simpler to more difficult subjects Similarly the early chapters are easier than the later ones making the book ideal for independent study Subjects begin with the laws of thermodynamics and statistical theory of information and of ensembles advancing to the ideal classical gases of polyatomic molecules non electrolyte liquids and solutions and surfaces Subsequent chapters explore imperfect classical and quantum gas phase transitions cooperative phenomena Green function methods the plasma transport in gases and metals Nyquist s theorem and its generalizations stochastic methods and many other topics **McGraw-Hill Education 500**

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Problems in Applied Thermodynamics Colin Bodsworth, A. S. Appleton, 1965 Problems and Solutions in General
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