

Introduction

This exam is a closed book, closed notes, no calculator, and no internet access. It is designed to test your understanding of the basic principles of chemistry, including atomic structure, chemical bonding, and stoichiometry. The exam is divided into two sections: Part A (Multiple Choice) and Part B (Short Answer). The total time for the exam is 1 hour and 30 minutes. The questions are of varying difficulty, and you are encouraged to attempt all questions. The exam is a closed book, closed notes, no calculator, and no internet access.

Instructions: Read the questions carefully and answer them in the spaces provided. Show your working for all calculations. The exam is a closed book, closed notes, no calculator, and no internet access.

Part A: Multiple Choice Questions



1. Which of the following is a physical property of a substance?
 - a. It can be broken into smaller pieces.
 - b. It has a fixed boiling point.
 - c. It reacts with oxygen to form a gas.
 - d. It has a characteristic color.
2. The chemical formula for water is H_2O . How many hydrogen atoms are present in one molecule of water?
 - a. 1
 - b. 2
 - c. 3
 - d. 4
3. The chemical formula for sodium chloride is $NaCl$. How many chlorine atoms are present in one molecule of sodium chloride?
 - a. 1
 - b. 2
 - c. 3
 - d. 4



4. Which of the following is a chemical property of a substance?
 - a. It has a fixed boiling point.
 - b. It reacts with oxygen to form a gas.
 - c. It has a characteristic color.
 - d. It can be broken into smaller pieces.
5. Complete the following table by filling in the missing values.

| Element | Atomic Number | Mass Number | Number of Protons | Number of Electrons | Number of Neutrons |
|---------|---------------|-------------|-------------------|---------------------|--------------------|
| H | 1 | 1 | 1 | 1 | 0 |
| He | 2 | 4 | 2 | 2 | 2 |
| Li | 3 | 7 | 3 | 3 | 4 |
| Be | 4 | 9 | 4 | 4 | 5 |
6. Calculate the relative atomic mass (A_r) of an element with the following isotopes. The relative atomic mass is defined as the weighted average of the masses of the isotopes.

| | | |
|-----------|-------|------|
| Isotope 1 | 20.0% | 20.0 |
| Isotope 2 | 80.0% | 22.0 |



Part B: Short Answer Questions



1. Write the chemical formula for the following compounds:
 - a. Sodium chloride
 - b. Magnesium sulfate
 - c. Calcium hydroxide
 - d. Ammonium nitrate
2. Calculate the relative atomic mass (A_r) of an element with the following isotopes. The relative atomic mass is defined as the weighted average of the masses of the isotopes.

| | | |
|-----------|-------|------|
| Isotope 1 | 25.0% | 25.0 |
| Isotope 2 | 75.0% | 27.0 |

10. Complete the following table by filling in the missing values.

| Element | Atomic Number | Mass Number | Number of Protons | Number of Electrons | Number of Neutrons |
|---------|---------------|-------------|-------------------|---------------------|--------------------|
| H | 1 | 1 | 1 | 1 | 0 |
| He | 2 | 4 | 2 | 2 | 2 |
| Li | 3 | 7 | 3 | 3 | 4 |
| Be | 4 | 9 | 4 | 4 | 5 |

11. Write the chemical formula for the following compounds: _____

Stoichiometry Phet Lab Answers

JA Banks



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Table of Contents Stoichiometry Phet Lab Answers

1. Understanding the eBook Stoichiometry Phet Lab Answers
 - The Rise of Digital Reading Stoichiometry Phet Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Phet Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry Phet Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Phet Lab Answers
 - Personalized Recommendations
 - Stoichiometry Phet Lab Answers User Reviews and Ratings
 - Stoichiometry Phet Lab Answers and Bestseller Lists
5. Accessing Stoichiometry Phet Lab Answers Free and Paid eBooks

- Stoichiometry Phet Lab Answers Public Domain eBooks
 - Stoichiometry Phet Lab Answers eBook Subscription Services
 - Stoichiometry Phet Lab Answers Budget-Friendly Options
6. Navigating Stoichiometry Phet Lab Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Stoichiometry Phet Lab Answers Compatibility with Devices
 - Stoichiometry Phet Lab Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Phet Lab Answers
 - Highlighting and Note-Taking Stoichiometry Phet Lab Answers
 - Interactive Elements Stoichiometry Phet Lab Answers
 8. Staying Engaged with Stoichiometry Phet Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Phet Lab Answers
 9. Balancing eBooks and Physical Books Stoichiometry Phet Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry Phet Lab Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Stoichiometry Phet Lab Answers
 - Setting Reading Goals Stoichiometry Phet Lab Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Stoichiometry Phet Lab Answers
 - Fact-Checking eBook Content of Stoichiometry Phet Lab Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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