

Introduction 11

An Electron's Address

Name: _____
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Introduction: Energy Levels and Sublevels

In addition to the energy level model, Bohr proposed the following model of energy levels. The energy levels within each one's energy level are the sublevels.

Sublevels are located inside energy levels just like sublevels are located inside each level. Each sublevel contains one or more orbitals and sublevels are:

TABLE 1

Energy Level	Number of sublevels in each energy level
1	1
2	2
3	3
4	4

Note that there is one orbital in each of the 1st sublevel, two orbitals in the 2nd sublevel, three orbitals in each of sublevel's orbitals in the 3rd energy level.

Critical Thinking Questions

- How many sublevels exist in the 1st energy level?
 The sublevel in the 1st energy level is the 1st energy level is the only sublevel.
- How many sublevels exist in the 2nd energy level?
 Two sublevels exist in the 2nd energy level in the 2nd energy level and the 1st sublevel.
- How many sublevels exist in the 3rd energy level?
 Three sublevels exist in the 3rd energy level and the 1st, 2nd, and 3rd sublevels.
- How many sublevels would you expect to exist in the 4th energy level?
 The number of sublevels in each energy level would be 4, 3, 2, and 1.
- How many sublevels would you expect to exist in the 5th energy level?
 The number of sublevels in each energy level would be 5, 4, 3, 2, and 1.
- How many sublevels would you expect to exist in the 6th energy level?
 The number of sublevels in each energy level would be 6, 5, 4, 3, 2, and 1.

Chemquest 11 Answers

Christian G. Meyer



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