



Chapter Test Answers

CHAPTER 1: CHEMICAL REACTIONS AND EQUATIONS

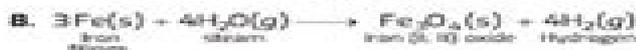
- (c) 4 and 2
- (d) double displacement reaction
- (a) only (i)
- (b) H_2 , CuO
- (a) Both Assertion (A) and Reason (R) are true and Reason (R) is the correct explanation of Assertion (A).
- (c) Assertion (A) is true but Reason (R) is false.

7. (i) When zinc metal is added to a solution of copper sulphate, a chemical reaction occurs where zinc displaces copper from the copper sulphate solution, forming zinc sulphate and solid copper metal. As a result, the blue colour of the copper sulphate solution fades, and a reddish-brown deposit of copper metal forms on the zinc.

(ii) When magnesium ribbon reacts with dilute hydrochloric acid, magnesium chloride salt will be formed and hydrogen gas is liberated. The reaction is exothermic, so the solution becomes warmer.

(iii) Yes, a displacement reaction is also a redox reaction because it involves the transfer of electrons where one substance is oxidised and another is reduced.

(iv) Galvanisation, electroplating and extraction of metals are some applications of displacement reactions in everyday life or industry.



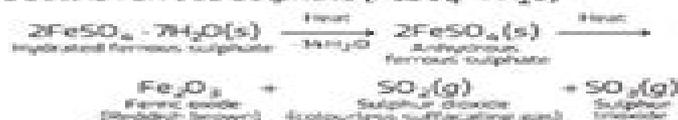
9. The given reaction is a redox reaction.

Oxidising agent – Fe_2O_3

Reducing agent – Al

- Decomposition reaction
- Displacement reaction
- Double displacement reaction
- Combination Reaction

11. The salt is ferrous sulphate ($FeSO_4 \cdot 7H_2O$)



12. Exothermic reaction is the reaction in which heat is released.

Endothermic reaction is the reaction in which heat is absorbed.

(i) Heating of coal in air to form carbon dioxide → Exothermic reaction.

(ii) Heating of limestone in a lime kiln to form quick lime → Endothermic reaction.

13. (i) The reaction will take place in presence of sunlight.

This reaction is used in black and white photography.

(ii) (a) Displacement reaction because Cl_2 is displacing I_2 from KI solution.

(b) Combination reaction because K combines with Cl_2 to form a single product, KCl.

14.

	Reaction	Substance	
		Oxidised	reduced
(i)	$\text{ZnO}_{(s)} + \text{C}_{(s)} \xrightarrow{\text{Oxidised}} \text{Zn}_{(s)} + \text{CO}_{(g)}$ <p style="text-align: center;"> reduced </p>	C	ZnO
(ii)	$4\text{Na}_{(s)} + \text{O}_{2(g)} \xrightarrow{\text{Oxidised}} 2\text{Na}_2\text{O}_{(s)}$ <p style="text-align: center;"> reduced </p>	Na	O_2
(iii)	$\text{CuO}_{(s)} + \text{H}_2(g) \xrightarrow{\text{reduced}} \text{Cu}_{(s)} + \text{H}_2\text{O}(l)$ <p style="text-align: center;"> Oxidised </p>	H_2	CuO

15. (i) X-copper (Cu). Process - Corrosion

(ii) Green coating – Basic copper carbonate [$CuCO_3 \cdot Cu(OH)_2$]

(iii) Painting and oiling.

16. (i) **Decomposition reaction:** Reaction in which a compound splits up into two or more simpler substances.

(ii) **Oxidation reaction:** The reaction in which oxygen is added to a substance or hydrogen is removed from a substance.

(iii) **Displacement reaction:** Reaction in which a more reactive element displaces a less reactive element from its compound.

(iv) **Displacement reaction:** Reaction in which a more reactive element displaces a less reactive element from its compound.

(v) **Combination reaction:** Reaction in which two or more substances combine to form a single substance.

17. The process in which new substances with new properties are formed by the rearrangement of atoms is known as chemical reaction.

Four observations which helps us to determine that a chemical reaction has taken place are as follows:

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Nitrogen and Sulfur Multiple Choice Questions 89 MCQs Organic and Nitrogen Compounds Multiple Choice Questions 54 MCQs Periodicity Multiple Choice Questions 202 MCQs Polymerization Multiple Choice Questions 36 MCQs Rates of Reaction Multiple Choice Questions 39 MCQs Reaction Kinetics Multiple Choice Questions 52 MCQs Redox Reactions and Electrolysis Multiple Choice Questions 55 MCQs States of Matter Multiple Choice Questions 66 MCQs Transition Elements Multiple Choice Questions 29 MCQs The chapter Alcohols and Esters MCQs covers topics of introduction to alcohols and alcohols reactions The chapter Atomic Structure and Theory MCQs covers topics of atom facts elements and atoms number of nucleons protons electrons and neutrons The chapter Benzene Chemical Compound MCQs covers topics of benzene arenes reaction phenol properties and reactions of phenol The chapter Carbonyl Compounds MCQs covers topics of carbonyl compounds aldehydes and ketone testing nucleophilic addition with HCN preparation of aldehydes and ketone reduction of aldehydes and ketone

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Table of Contents Chemistrymc Com Chapter Test Answers

1. Understanding the eBook Chemistrymc Com Chapter Test Answers
 - The Rise of Digital Reading Chemistrymc Com Chapter Test Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistrymc Com Chapter Test Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistrymc Com Chapter Test Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistrymc Com Chapter Test Answers
 - Personalized Recommendations
 - Chemistrymc Com Chapter Test Answers User Reviews and Ratings
 - Chemistrymc Com Chapter Test Answers and Bestseller Lists

5. Accessing Chemistrymc Com Chapter Test Answers Free and Paid eBooks
 - Chemistrymc Com Chapter Test Answers Public Domain eBooks
 - Chemistrymc Com Chapter Test Answers eBook Subscription Services
 - Chemistrymc Com Chapter Test Answers Budget-Friendly Options
6. Navigating Chemistrymc Com Chapter Test Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistrymc Com Chapter Test Answers Compatibility with Devices
 - Chemistrymc Com Chapter Test Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistrymc Com Chapter Test Answers
 - Highlighting and Note-Taking Chemistrymc Com Chapter Test Answers
 - Interactive Elements Chemistrymc Com Chapter Test Answers
8. Staying Engaged with Chemistrymc Com Chapter Test Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistrymc Com Chapter Test Answers
9. Balancing eBooks and Physical Books Chemistrymc Com Chapter Test Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistrymc Com Chapter Test Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistrymc Com Chapter Test Answers
 - Setting Reading Goals Chemistrymc Com Chapter Test Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistrymc Com Chapter Test Answers
 - Fact-Checking eBook Content of Chemistrymc Com Chapter Test Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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