

Atomic structure

Label the following diagram with the **correct** proton, neutron, electron and nucleus.



Carbon atom

2. Now complete the table below, identifying the charges, mass and place found for each of the sub-atomic particles.

Sub-atomic particle	Place found	Mass	Charge
proton			
neutron			
electron			

3. Why are atoms neutral overall?

.....

.....

Atomic structure 2

Use the information in the periodic table to help you complete the table below:

Element	Protons	Neutrons	Electrons
Carbon			
Sodium			
Magnesium			
Isotope			
Bromine			
Hydrogen			
Aluminium			
Argon			

3. What is an isotope?

.....

.....

Calculating Relative Atomic Mass "A_r"

Calculate the relative atomic mass of the following elements, given their mass and abundance.

a) Bromine, consisting of 50% Br-79 and 50% Br-81

.....

b) Copper, consisting of 69.0% Cu-63 and 30.9% Cu-65

.....

c) Silver, consisting of 97.5% Ag-107 and 2.5% Ag-109

.....

The Periodic Table

1. What element is found in:

- a) Group 2 Period 4 _____
- b) Group 1 Period 1 _____
- c) Group 7 Period 2 _____
- d) Group 8 Period 2 _____

2. What group and period are the following elements found in?

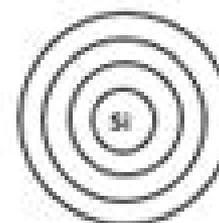
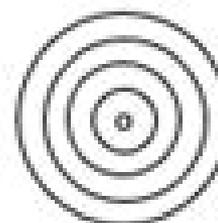
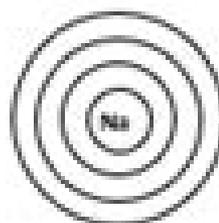
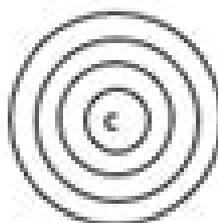
- a) Lithium _____
- b) Phosphorus _____
- c) Iron _____
- d) Rubidium _____

Electron structure

1. How many electrons can each shell hold?

- a) 1st shell _____
- b) 2nd shell _____
- c) 3rd shell _____
- d) 4th shell _____

2. Complete the electron structure diagrams for the following elements by drawing electrons on the shells. Under each diagram, WRITE the electron configuration, e.g. 2, 8, 1



The Periodic Table

1. Shade the metals in blue
2. Shade the non-metals in red
3. Shade the metalloids in green
4. What is special about the properties of metalloids? _____

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 - The Rise of Digital Reading Chemistry End Off Topic Assessment Unit C2
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry End Off Topic Assessment Unit C2
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry End Off Topic Assessment Unit C2
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry End Off Topic Assessment Unit C2
 - Personalized Recommendations
 - Chemistry End Off Topic Assessment Unit C2 User Reviews and Ratings
 - Chemistry End Off Topic Assessment Unit C2 and Bestseller Lists
5. Accessing Chemistry End Off Topic Assessment Unit C2 Free and Paid eBooks

- Chemistry End Off Topic Assessment Unit C2 Public Domain eBooks
 - Chemistry End Off Topic Assessment Unit C2 eBook Subscription Services
 - Chemistry End Off Topic Assessment Unit C2 Budget-Friendly Options
6. Navigating Chemistry End Off Topic Assessment Unit C2 eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry End Off Topic Assessment Unit C2 Compatibility with Devices
 - Chemistry End Off Topic Assessment Unit C2 Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry End Off Topic Assessment Unit C2
 - Highlighting and Note-Taking Chemistry End Off Topic Assessment Unit C2
 - Interactive Elements Chemistry End Off Topic Assessment Unit C2
 8. Staying Engaged with Chemistry End Off Topic Assessment Unit C2
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry End Off Topic Assessment Unit C2
 9. Balancing eBooks and Physical Books Chemistry End Off Topic Assessment Unit C2
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry End Off Topic Assessment Unit C2
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistry End Off Topic Assessment Unit C2
 - Setting Reading Goals Chemistry End Off Topic Assessment Unit C2
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistry End Off Topic Assessment Unit C2
 - Fact-Checking eBook Content of Chemistry End Off Topic Assessment Unit C2
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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