

## Chapter 6 Practice Test: Chemical Bonding

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- \_\_\_\_\_ 1. The charge on an ion is  
a. always positive.  
b. always negative.  
c. either positive or negative.  
d. zero.
- \_\_\_\_\_ 2. According to the octet rule, a calcium atom has a tendency to  
a. lose one electron.  
b. lose two electrons.  
c. gain one electron.  
d. gain two electrons.
- \_\_\_\_\_ 3. If a compound forms by ionic bonding, which is *not* true?  
a. A positively charged atom or group of atoms attracts a negatively charged atom or group of atoms.  
b. The net charge of the compound is zero.  
c. The compound contains just two atoms, each of opposite charge.  
d. Several ions group together in a tightly packed structure.
- \_\_\_\_\_ 4. The only property listed that is *not* characteristic of ionic compounds is  
a. high melting point.  
b. hardness.  
c. lack of crystal structure.  
d. good conductor of electricity.
- \_\_\_\_\_ 5. Which formula listed below represents a polyatomic ion?  
a.  $\text{NH}_4^+$   
b.  $\text{H}_2\text{SO}_4$   
c.  $\text{NaCl}$   
d.  $\text{H}_2\text{O}$
- \_\_\_\_\_ 6. The melting points of ionic compounds are higher than the melting points of molecular compounds because  
a. ionic substances tend to vaporize at room temperature.  
b. ionic substances are brittle.  
c. attractive forces between ions are greater than the attractive forces between molecules.  
d. the numbers of positive and negative charges are equal in an ionic compound.

# Chemistry Chemical Bonding Test Answers

**Leo J. Malone**



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## **Table of Contents Chemistry Chemical Bonding Test Answers**

1. Understanding the eBook Chemistry Chemical Bonding Test Answers
  - The Rise of Digital Reading Chemistry Chemical Bonding Test Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chemical Bonding Test Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Chemical Bonding Test Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chemical Bonding Test Answers
  - Personalized Recommendations
  - Chemistry Chemical Bonding Test Answers User Reviews and Ratings

- Chemistry Chemical Bonding Test Answers and Bestseller Lists
- 5. Accessing Chemistry Chemical Bonding Test Answers Free and Paid eBooks
  - Chemistry Chemical Bonding Test Answers Public Domain eBooks
  - Chemistry Chemical Bonding Test Answers eBook Subscription Services
  - Chemistry Chemical Bonding Test Answers Budget-Friendly Options
- 6. Navigating Chemistry Chemical Bonding Test Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Chemical Bonding Test Answers Compatibility with Devices
  - Chemistry Chemical Bonding Test Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Chemical Bonding Test Answers
  - Highlighting and Note-Taking Chemistry Chemical Bonding Test Answers
  - Interactive Elements Chemistry Chemical Bonding Test Answers
- 8. Staying Engaged with Chemistry Chemical Bonding Test Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Chemical Bonding Test Answers
- 9. Balancing eBooks and Physical Books Chemistry Chemical Bonding Test Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Chemical Bonding Test Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Chemical Bonding Test Answers
  - Setting Reading Goals Chemistry Chemical Bonding Test Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Chemical Bonding Test Answers
  - Fact-Checking eBook Content of Chemistry Chemical Bonding Test Answers
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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