

Unit Two Chemistry Exam

Multiple Choice 20 marks

You need to record your answers in the appropriate section in the ANSWER BOOKLET.

Question 1

When ethanol dissolves in water the bonds formed between the ethanol and water molecules are known as

- A. dispersion forces
- B. ion-dipole attraction
- C. hydrogen bonds
- D. ionisation

Question 2

When NaOH dissociates in water, the equation that describes the process is

- A. $\text{NaOH (s)} \rightarrow \text{Na}^+ \text{(aq)} + \text{OH}^- \text{(aq)}$
- B. $\text{NaOH (s)} + \text{H}_2\text{O (l)} \rightarrow \text{NaO}^- \text{(aq)} + \text{H}_2\text{O}^+ \text{(aq)}$
- C. $\text{NaOH (s)} + \text{H}_2\text{O (l)} \rightarrow \text{NaOH}_2^+ \text{(aq)} + \text{OH}^- \text{(aq)}$
- D. $\text{NaOH (aq)} + \text{HCl (aq)} \rightarrow \text{NaCl (s)} + \text{H}_2\text{O (l)}$

Question 3

How much heat energy is required to raise the temperature of 250 mL of water from 17 °C to 100 °C? Assume that 1 mL of water has a mass of 1 g and that the specific heat capacity of water is $4.2 \text{ J g}^{-1} \text{ }^\circ\text{C}^{-1}$.

- A. 8.7 J
- B. 87,150 KJ
- C. 8.7 KJ
- D. $8.7 \times 10^4 \text{ J}$

Question 4

A solution containing 0.25 g of NaCl dissolved in 50 mL of deionised water will have a concentration of

- A. 5.0 gL^{-1}
- B. 4.3 M
- C. $5.0 \times 10^{-3} \text{ mgL}^{-1}$
- D. $2.5 \times 10^{-4} \text{ molL}$

Question 5

A solution of CuSO_4 has a concentration of 2.50 M. This concentration is equivalent to

- A. 399 %m/v
- B. 39.9 %m/v
- C. 0.250 %m/v
- D. 25.0 m/v

Chemistry Unit 2 Test Answers

KJ Lindholm-Leary



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Table of Contents Chemistry Unit 2 Test Answers

1. Understanding the eBook Chemistry Unit 2 Test Answers
 - The Rise of Digital Reading Chemistry Unit 2 Test Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Unit 2 Test Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Unit 2 Test Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Unit 2 Test Answers
 - Personalized Recommendations
 - Chemistry Unit 2 Test Answers User Reviews and Ratings
 - Chemistry Unit 2 Test Answers and Bestseller Lists
5. Accessing Chemistry Unit 2 Test Answers Free and Paid eBooks
 - Chemistry Unit 2 Test Answers Public Domain eBooks
 - Chemistry Unit 2 Test Answers eBook Subscription Services
 - Chemistry Unit 2 Test Answers Budget-Friendly Options

6. Navigating Chemistry Unit 2 Test Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Unit 2 Test Answers Compatibility with Devices
 - Chemistry Unit 2 Test Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Unit 2 Test Answers
 - Highlighting and Note-Taking Chemistry Unit 2 Test Answers
 - Interactive Elements Chemistry Unit 2 Test Answers
8. Staying Engaged with Chemistry Unit 2 Test Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Unit 2 Test Answers
9. Balancing eBooks and Physical Books Chemistry Unit 2 Test Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Unit 2 Test Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Unit 2 Test Answers
 - Setting Reading Goals Chemistry Unit 2 Test Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Unit 2 Test Answers
 - Fact-Checking eBook Content of Chemistry Unit 2 Test Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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