

# CHEMISTRY

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## PERIODIC TABLE OF THE ELEMENTS

The periodic table is color-coded by groups:
 

- 1 (pink):** s-subshell, 1 orbital
- 2 (orange):** s-subshell, 1 orbital
- 13 (light blue):** s-subshell, 1 orbital; p-subshell, 3 orbitals
- 14 (medium blue):** s-subshell, 1 orbital; p-subshell, 3 orbitals
- 15 (dark blue):** s-subshell, 1 orbital; p-subshell, 3 orbitals
- 16 (teal):** s-subshell, 1 orbital; p-subshell, 3 orbitals
- 17 (green):** s-subshell, 1 orbital; p-subshell, 3 orbitals
- 18 (light green):** s-subshell, 1 orbital; p-subshell, 3 orbitals
- 21 (purple):** s-subshell, 1 orbital; d-subshell, 5 orbitals
- 22 (blue-purple):** s-subshell, 1 orbital; d-subshell, 5 orbitals
- 23 (blue):** s-subshell, 1 orbital; d-subshell, 5 orbitals
- 24 (teal):** s-subshell, 1 orbital; d-subshell, 5 orbitals
- 25 (green):** s-subshell, 1 orbital; d-subshell, 5 orbitals
- 26 (light green):** s-subshell, 1 orbital; d-subshell, 5 orbitals
- 27 (yellow-green):** s-subshell, 1 orbital; d-subshell, 5 orbitals
- 28 (yellow):** s-subshell, 1 orbital; d-subshell, 5 orbitals
- 29 (orange):** s-subshell, 1 orbital; d-subshell, 5 orbitals
- 30 (light orange):** s-subshell, 1 orbital; d-subshell, 5 orbitals
- 31 (pink):** s-subshell, 1 orbital; p-subshell, 3 orbitals
- 32 (light pink):** s-subshell, 1 orbital; p-subshell, 3 orbitals



### ATOMIC STRUCTURE

Atomic Number,  $Z$ , is the number of protons in the nucleus. For neutral atoms,  $Z = +$  the number of electrons. The mass number,  $A$ , is the sum of protons and neutrons in the nucleus. Isotopes of an element have the same  $Z$ , different  $A$ . The atomic mass is the weighted average of isotopes. The atomic weight is the weighted average of isotopes for a given element.

### ATOMIC QUANTUM NUMBERS AND ORBITALS

Principal quantum number,  $n$ , is the number of shells.  $n = 1, 2, 3, \dots$ . Subshell quantum numbers are  $l = 0, 1, 2, \dots, n-1$ . Magnetic quantum number,  $m_l$ , is the number of orbitals in a subshell. Spin quantum number,  $m_s$ , is the number of electrons in a subshell.

### NUCLEAR CHEMISTRY

Nuclear reactions occur when nucleons (protons and neutrons) are added or removed. Energy is released or absorbed. The mass defect is the difference between the mass of the nucleus and the sum of the masses of its constituent nucleons.



### MANY-ELECTRON ATOMS

Orbitals and quantum numbers are used to describe the behavior of electrons in many-electron atoms. The Pauli Exclusion Principle states that no two electrons can have the same set of quantum numbers.

n	l	m <sub>l</sub>	m <sub>s</sub>	Orbitals
1	0	0	±1/2	1 orbital
2	0	0	±1/2	1 orbital
2	1	0, ±1	±1/2	3 orbitals
3	0	0	±1/2	1 orbital
3	1	0, ±1	±1/2	3 orbitals
3	2	0, ±1, ±2	±1/2	5 orbitals



Atomic configuration: The arrangement of electrons in an atom. The Pauli Exclusion Principle states that no two electrons can have the same set of quantum numbers. Hund's Rule states that electrons will occupy orbitals singly before pairing up.

# Chemistry Quickstudy Reference Guides Academic

**Jin-Ying Zhang**



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